

1 Mole to Mole Relationship between Cu and

Ag_Essay1_Experimental Data Table

Record the following data in the table below.

mass of empty beaker	110.000 g
mass of beaker with silver nitrate (after all additions) and copper	331.634 g
mass of beaker with silver	113.395 g

2 Mole to Mole Relationship between Cu and

Ag_Essay2_Moles Silver Nitrate Added to Reaction

How many moles of silver nitrate were added to the reaction?

$$n = M \times V$$

$$n = 0.2 \text{ M} \times 0.200 \text{ L} = 0.04 \text{ mol AgNO}_3$$

3 Mole to Mole Relationship between Cu and

Ag_Essay3_Silver Ions Consumed

Did all the silver ions get consumed in the reaction? The molar mass of silver is 107.87 g/mol. Justify your answer.

$$n_{\text{Ag}^+} = M \times V$$

$$n_{\text{Ag}^+ \text{ added}} = 0.2 \text{ M} \times 0.200 \text{ L} = 0.04 \text{ mol Ag}^+$$

$$n_{\text{Ag}^+ \text{ consumed}} = \frac{m}{MM} = \frac{113.395 \text{ g} - 110.000 \text{ g}}{107.87 \text{ g/mol}} = 0.03147 \text{ mol Ag}$$

No, not all the silver ions were consumed in the reaction.

4 Mole to Mole Relationship between Cu and

Ag_Essay4_Silver Nitrate Plus Zinc Reaction

Mole Relationships Answers

Kathleen Armour

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Mole Relationships Answers:

Unveiling the Power of Verbal Artistry: An Mental Sojourn through **Mole Relationships Answers**

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