MOLARITY BY DILUTION

79+52 = 162 Name _____

Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid = $M \times V$ then.

M, $\times V_{\rm c} = M_{\rm c} \times V_{\rm c}$. Solve the following problems.

 How much concentrated 18 M sulfuric acid is needed to prepare 250 mL of a 6.0 M solution?

83 mL]

 How much concentrated 12 M hydrochloric acid is needed to prepare 100,mL of a 2.0 M solution?

- 17mL/
- To what volume should 25 mL of 15 M nitric acid be diluted to prepare a 3.0 M solution?

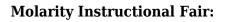
4. To how much water should 50, mL of 12 M hydrochloric acid be added to produce a 4.0 M solution?

100ml H20)

5. To how much water should 100. mL of 18 M sulfuric acid be added to prepare a 1.5 M solution?

Molarity Instructional Fair

Ensheng Dong



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Table of Contents Molarity Instructional Fair

- 1. Understanding the eBook Molarity Instructional Fair
 - The Rise of Digital Reading Molarity Instructional Fair
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Molarity Instructional Fair
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Instructional Fair
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Molarity Instructional Fair
 - Personalized Recommendations
 - Molarity Instructional Fair User Reviews and Ratings
 - Molarity Instructional Fair and Bestseller Lists
- 5. Accessing Molarity Instructional Fair Free and Paid eBooks
 - Molarity Instructional Fair Public Domain eBooks
 - Molarity Instructional Fair eBook Subscription Services
 - Molarity Instructional Fair Budget-Friendly Options
- 6. Navigating Molarity Instructional Fair eBook Formats

- ∘ ePub, PDF, MOBI, and More
- Molarity Instructional Fair Compatibility with Devices
- Molarity Instructional Fair Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Molarity Instructional Fair
 - Highlighting and Note-Taking Molarity Instructional Fair
 - Interactive Elements Molarity Instructional Fair
- 8. Staying Engaged with Molarity Instructional Fair
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Instructional Fair
- 9. Balancing eBooks and Physical Books Molarity Instructional Fair
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Instructional Fair
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Molarity Instructional Fair
 - Setting Reading Goals Molarity Instructional Fair
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molarity Instructional Fair
 - Fact-Checking eBook Content of Molarity Instructional Fair
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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