

POLARITY REVIEW

1. Molarity is a type of _____, similar to g/L , mg/mL , cm^3/cm^3 , etc. etc. etc.
2. Molarity is defined as M of _____ dissolved within one _____ of solution.
3. M is the unit for molarity, but the unit for molarity can also be written as:
4. Analyze the images below and identify the molarity of the solutions pictured.

2 moles NaCl dissolved in 2 liters of solution



_____ M

4.8 moles NaCl dissolved in 4.8 liters of solution



_____ M

4 moles NaCl dissolved in 8.8 liters of solution



_____ M

4.8 moles NaCl dissolved in 8.8 liters of solution



_____ M

2 moles NaCl dissolved in 4.8 liters of solution



_____ M

11 moles NaCl dissolved in 1.1 liters of solution



_____ M

5. Analyze the images below and identify how many total moles of the compound are present

1.5 L of 0.5M solution



_____ moles

1000 mL of 1.00M solution



_____ moles

2000 mL of 0.25M solution



_____ moles

0.5L of 1.00M solution



_____ moles

2000 mL of 0M solution



_____ moles

0.001L of 100M solution



_____ moles

6. Analyze the images below and answer the prompts accordingly. Be sure to include the correct units and identity of the solution on question. Ex: 2 moles CH_4 or 3M MgO

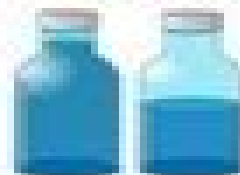
How many moles of NaCl are in 1000 mL of a 0.50 M solution?



What is the concentration in molarity of 2000 mL of a solution containing 10 moles of NaCl ?



How many moles of NaCl are in 1.5 L of a 0.50 M solution?



Answer the below 5a and 5b with the number of grams. 5a. How many grams of sugar are in 100 mL of a 10% solution of sugar? 5b. How many grams of sugar are in 100 mL of a 10% solution of sugar?



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Kaplan Test Prep



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