

Chapter 15

Oxidation-Reduction Reactions

SECTION 1

Transfer electrons around the following equations for the species provided.

1. _____ Add the following species into separate redox reactions.
a. $\text{Cu}^{2+} + \text{Mg} \rightarrow \text{Cu} + \text{Mg}^{2+}$
b. $\text{Fe}^{3+} + \text{Fe} \rightarrow \text{Fe}^{2+}$
c. $\text{Mg}^{2+} + \text{Zn} \rightarrow \text{Zn}^{2+}$

Answers:

2. Assign each element its oxidation number in the indicated species as they follow:

- a. _____ in Fe_2O_3 in Fe_2O_3
b. _____ in H_2SO_4 in H_2SO_4
c. _____ in CO_2 in CO_2
d. _____ in H_2O in H_2O
e. _____ in HCl in HCl
f. _____ in FeSO_4 in FeSO_4
g. _____ in FeSO_4 in FeSO_4

3. In each of the following half-reactions, determine the electron:

- a. _____ in $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
b. _____ in $\text{Zn}^{2+} \rightarrow \text{Zn} + 2e^-$
c. _____ in $\text{Cu}^{2+} + 2e^- \rightarrow \text{Cu}$

- Answers:** a. _____ of electrons in that during half-reactions represent
oxidation process.

4. Name compounds other than those listed in Figure 15.1 of the text, but the following:

Answers:

- a. _____ is composed of a mixture of 50% of
oxidation state.
b. _____ is a gas.
c. _____ is a gas, which is a mixture of 50% of
oxidation state.
d. _____ is a gas, which is a mixture of 50% of
oxidation state.

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David Baud



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Table of Contents Modern Chemistry Chapter 19 Answers

1. Understanding the eBook Modern Chemistry Chapter 19 Answers
 - The Rise of Digital Reading Modern Chemistry Chapter 19 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 19 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Chapter 19 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter 19 Answers
 - Personalized Recommendations
 - Modern Chemistry Chapter 19 Answers User Reviews and Ratings
 - Modern Chemistry Chapter 19 Answers and Bestseller Lists

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 - Modern Chemistry Chapter 19 Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Modern Chemistry Chapter 19 Answers
 - Interactive Elements Modern Chemistry Chapter 19 Answers
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Chapter 19 Answers
9. Balancing eBooks and Physical Books Modern Chemistry Chapter 19 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Chapter 19 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Chapter 19 Answers
 - Setting Reading Goals Modern Chemistry Chapter 19 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Chapter 19 Answers
 - Fact-Checking eBook Content of Modern Chemistry Chapter 19 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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