

## Practice Final Exam Chem 112

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. Three hundred milliliters (150 mL) of 4.50 M sugar solution is diluted to 600 mL by the addition of pure water. Which of these statements about the solution is correct?  
a. The moles of water remain the same.    b. The moles of sugar decrease.    c. The concentration of sugar increases to 5.0 M.    d. The moles of sugar remain constant.    e. All of these statements are incorrect.
- \_\_\_\_\_ 2. How many moles of solute are present in 75.0 grams of solution that is 1.30% by mass glucose ( $C_6H_{12}O_6$ )?  $MM(C_6H_{12}O_6) = 180.16 \text{ g/mol}$   
a.  $5.4 \times 10^{-3}$  moles    b. 0.800 moles    c. 176 moles    d.  $1.76 \times 10^{-3}$  moles    e. 225 moles
- \_\_\_\_\_ 3. A concentrated reagent is sold as a solution that is 60.0 mass % solute. If the molar mass of the solute is 244 g/mol and the density of this solution is 1.20 g/mL, what is the molarity of the solution?  
a. 10.0 M    b. 2.95 M    c. 15.5 M    d. 0.72 M    e. 1.45 M
- \_\_\_\_\_ 4. The boiling point of pure benzene ( $C_6H_6$ ) is 80.10 °C. A solution of 14.2 g of a solid dissolved in 64.0 g of benzene boils at 81.09 °C. If  $k_b$  for benzene is 2.53 °C/molal, the molar mass of the solid is:  
a. 662 g/mol    b. 56.0 g/mol    c. 92.0 g/mol    d. 628 g/mol    e. 568 g/mol
- \_\_\_\_\_ 5. What is the freezing point for a 0.71 m solution of glycerol ( $C_3H_8O_3$ ) in ethanol? Freezing point(ethanol) = -114.6 °C;  $K_f(\text{ethanol}) = 1.99 \text{ °C/m}$   
a. 0.10 °C    b. -113.2 °C    c. -1.41 °C    d. -116.0 °C    e. none of these
- \_\_\_\_\_ 6. At 450 °C,  $K_c = 0.231$  for the reaction:



What is the value of the equilibrium constant for the following reaction at 450 °C?



- a. 0.481    b. 0.156    c. 0.231    d. 0.053    e. 4.329

- \_\_\_\_\_ 7. Consider the following system at equilibrium and its corresponding concentration equilibrium constant:



What is the value of  $K_p$  for this reaction at 470 °C?

- a.  $K_p = 3.80 \times 10^{-6}$     b.  $K_p = 8.53 \times 10^{-6}$     c.  $K_p = 4.32 \times 10^{-5}$     d.  $K_p = 1.10 \times 10^{-4}$     e.  $K_p = 1.19 \times 10^4$

- \_\_\_\_\_ 8. For the reaction  $\text{I}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2 \text{ICl}(\text{g})$ , the relation between  $K_p$  and  $K_c$  is:  
a.  $K_p = K_c$     b.  $K_p = K_c(RT)^2$     c.  $K_p = K_c(RT)^{-2}$     d.  $K_c = K_p(RT)^2$     e.  $K_p = K_c(RT)^4$

# Mc Chemistry Final Exam Answers

**Scott C. Dulebohn**



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## **Table of Contents Mc Chemistry Final Exam Answers**

1. Understanding the eBook Mc Chemistry Final Exam Answers
  - The Rise of Digital Reading Mc Chemistry Final Exam Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Mc Chemistry Final Exam Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mc Chemistry Final Exam Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mc Chemistry Final Exam Answers
  - Personalized Recommendations
  - Mc Chemistry Final Exam Answers User Reviews and Ratings
  - Mc Chemistry Final Exam Answers and Bestseller Lists
5. Accessing Mc Chemistry Final Exam Answers Free and Paid eBooks

- Mc Chemistry Final Exam Answers Public Domain eBooks
  - Mc Chemistry Final Exam Answers eBook Subscription Services
  - Mc Chemistry Final Exam Answers Budget-Friendly Options
6. Navigating Mc Chemistry Final Exam Answers eBook Formats
- ePub, PDF, MOBI, and More
  - Mc Chemistry Final Exam Answers Compatibility with Devices
  - Mc Chemistry Final Exam Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Mc Chemistry Final Exam Answers
  - Highlighting and Note-Taking Mc Chemistry Final Exam Answers
  - Interactive Elements Mc Chemistry Final Exam Answers
8. Staying Engaged with Mc Chemistry Final Exam Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mc Chemistry Final Exam Answers
9. Balancing eBooks and Physical Books Mc Chemistry Final Exam Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mc Chemistry Final Exam Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Mc Chemistry Final Exam Answers
- Setting Reading Goals Mc Chemistry Final Exam Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mc Chemistry Final Exam Answers
- Fact-Checking eBook Content of Mc Chemistry Final Exam Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

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- Interactive and Gamified eBooks

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