Math and Chemistry

In chemistry, quantities and values are often very, very large or very, very small. One mole of H₂O contains a huge number (602 214 199 000 000 000 000 000 000) of molecules. The mass of each molecule is minuscule (0.000 000 000 000 000 000 000 000 9g). To calculate the mass of one mole of H₂O using the above data, you multiply the number of molecules in a mole by the mass of an individual molecule. Imagine doing this by hand:

602 214 199 000 000 000 000 000

× 0.000 000 000 000 000 000 000 029 9 g 18.0062045591 g

There is another problem with writing all these zeros: It implies that they are all significant digits. Yet the nonfinal zeros to the right of the decimal point in the mass of an H₂O molecule are used as placeholders. They are not significant digits. There are only three significant digits in

0.000 000 000 000 000 000 000 000 025 9 g

Scientific Notation

To simplify reporting large numbers and doing calculations, you can use scientific notation. As you learned in an earlier mathematics course, scientific notation works by using powers of 10 as multipliers.

The number of particles in one mole of a substance (the Avegacire constant) is often rounded down to 622 000 000 000 000 000 000 000. When performing calculations, however, this number is still too unvertelely. To simplify it, you can express it as a number between 1 and 10, multiplied by factors of 10. To do this, move the decimal point behind the left-most non-zero digit, counting the number of places that the decimal point noves. The number of places moved is the exponent for the base 10.



The decimal point moves to the left.

Numbers less than 1 are converted into scientific notation by moving the decimal point to the right of the left-most non-zero digit. The exponent is expressed as a negative power of 10.

Floure 6.3 The decimal point moves to the right.

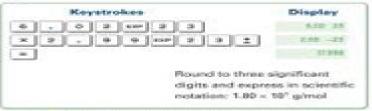
Once you have numbers in scientific notation, calculations become much simpler. To find the mass of one mole of H_2O , you can multiply $\{6.02 \times 10^{23} / \text{mol}\} \times \{2.99 \times 10^{-23} \text{ g}\}$. To do this without a calculator, you first multiply.

 $6.02 \times 2.99 = 17.9998 \rightarrow 18.0$ (rounded up to three significant digits)

When multiplying, exponents are added to find the new exponent: -23 + 23 = 0.

This gives 10.0×10^6 g/mol. In scientific notation, however, there must be only one digit before the decimal place. Here the decimal place must be moved one place to the left: 1.80×10^6 g/mol.

Figure E.3 shows how to do this calculation with a scientific calculator. When you enter an exponent on a scientific calculator, you do not have to enter [× 10] at any time.



repure £3 On some scientific calculators, the --- key is labelled ---. Key in negative exponents by entering the exponent, then striking the | | | | | | | | | | |

Rules for Scientific Notation

When multiplying, exponents are added algebraically.

$$(2.37 \times 10^{4} \text{ g}) \times (1.89 \times 10^{-1} \text{ g})$$

= $(2.37 \times 1.89) \times 10^{(4 - (-1))}$
= $4.4793 \times 10^{3} \text{ g} \rightarrow 4.48 \times 10^{3} \text{ g}$

 When dividing, exponents are subtracted algebraically.

 When adding or subtracting numbers in scientific notation, the numbers must be converted to the same power of 10 as the measurement with the greatest power of 10. Once the numbers are all expressed to the same power of 10, the power of 10 is neither added nor subtracted in the calculation.

$$(8.22 \times 10^4 \text{ cm}) + (5.66 \times 10^3 \text{ cm})$$

= $(8.22 \times 10^4 \text{ cm}) + (.566 \times 10^4 \text{ cm})$
= $8.786 \times 10^4 \text{ cm} \rightarrow 8.79 \times 10^4 \text{ cm}$

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Table of Contents Mathematics 11 Mcgraw Hill Ryerson Solutions

- 1. Understanding the eBook Mathematics 11 Mcgraw Hill Ryerson Solutions
 - The Rise of Digital Reading Mathematics 11 Mcgraw Hill Ryerson Solutions
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Mathematics 11 Mcgraw Hill Ryerson Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - \circ Features to Look for in an Mathematics 11 Mcgraw Hill Ryerson Solutions
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mathematics 11 Mcgraw Hill Ryerson Solutions
 - Personalized Recommendations
 - Mathematics 11 Mcgraw Hill Ryerson Solutions User Reviews and Ratings
 - Mathematics 11 Mcgraw Hill Ryerson Solutions and Bestseller Lists
- 5. Accessing Mathematics 11 Mcgraw Hill Ryerson Solutions Free and Paid eBooks

- Mathematics 11 Mcgraw Hill Ryerson Solutions Public Domain eBooks
- Mathematics 11 Mcgraw Hill Ryerson Solutions eBook Subscription Services
- Mathematics 11 Mcgraw Hill Ryerson Solutions Budget-Friendly Options
- 6. Navigating Mathematics 11 Mcgraw Hill Ryerson Solutions eBook Formats
 - o ePub, PDF, MOBI, and More
 - Mathematics 11 Mcgraw Hill Ryerson Solutions Compatibility with Devices
 - Mathematics 11 Mcgraw Hill Ryerson Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mathematics 11 Mcgraw Hill Ryerson Solutions
 - Highlighting and Note-Taking Mathematics 11 Mcgraw Hill Ryerson Solutions
 - Interactive Elements Mathematics 11 Mcgraw Hill Ryerson Solutions
- 8. Staying Engaged with Mathematics 11 Mcgraw Hill Ryerson Solutions
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mathematics 11 Mcgraw Hill Ryerson Solutions
- 9. Balancing eBooks and Physical Books Mathematics 11 Mcgraw Hill Ryerson Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mathematics 11 Mcgraw Hill Ryerson Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mathematics 11 Mcgraw Hill Ryerson Solutions
 - Setting Reading Goals Mathematics 11 Mcgraw Hill Ryerson Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mathematics 11 Mcgraw Hill Ryerson Solutions
 - Fact-Checking eBook Content of Mathematics 11 Mcgraw Hill Ryerson Solutions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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