

5

INTERPRETING GRAPHICS

Use with Section 5.3



Figure 1 The emission spectrum and orbit-transition diagram for hydrogen.

Table 1

Transition	$E(J)$	$r(m)$	$\lambda(m)$	Type of Radiation
$n=6 \rightarrow n=5$	2.66×10^{-20}	4.01×10^{10}	7.48×10^{-6}	Infrared
$n=6 \rightarrow n=4$	7.57×10^{-20}	1.14×10^{11}	2.63×10^{-5}	Infrared
$n=6 \rightarrow n=3$	1.82×10^{-19}	2.15×10^{11}	1.09×10^{-4}	Infrared
$n=6 \rightarrow n=2$	4.84×10^{-19}	7.30×10^{10}	4.11×10^{-5}	Infrared
$n=6 \rightarrow n=1$	2.12×10^{-18}	3.10×10^{10}	9.37×10^{-8}	Visible
$n=5 \rightarrow n=4$	4.91×10^{-20}	7.46×10^{10}	4.01×10^{-5}	Infrared
$n=5 \rightarrow n=3$	1.55×10^{-19}	2.34×10^{11}	1.28×10^{-4}	Infrared
$n=5 \rightarrow n=2$	4.56×10^{-19}	6.88×10^{10}	4.36×10^{-5}	Infrared
$n=5 \rightarrow n=1$	2.09×10^{-18}	3.15×10^{10}	9.52×10^{-8}	Visible
$n=4 \rightarrow n=3$	1.06×10^{-19}	1.60×10^{11}	1.87×10^{-4}	Infrared
$n=4 \rightarrow n=2$	4.09×10^{-19}	6.17×10^{10}	4.86×10^{-7}	Infrared
$n=4 \rightarrow n=1$	2.04×10^{-18}	3.08×10^{10}	9.72×10^{-8}	Visible
$n=3 \rightarrow n=2$	3.03×10^{-19}	4.58×10^{10}	6.53×10^{-7}	Infrared
$n=3 \rightarrow n=1$	1.94×10^{-18}	2.43×10^{10}	1.02×10^{-7}	Visible
$n=2 \rightarrow n=1$	1.64×10^{-18}	2.48×10^{10}	1.20×10^{-7}	Visible

$$\frac{2.18 \times 10^{-18} J}{6.626 \times 10^{-34} J \cdot s} = \frac{6.626 \times 10^{-34} J \cdot s}{h} \cdot \frac{1}{\lambda}$$

$$V = 9.01 \times 10^8$$

$$h = 6.626 \times 10^{-34}$$

$$c = 2.998 \times 10^8 \text{ m/s}$$

$$c = \lambda \nu$$

$$E = h \nu$$

$$2.998 \times 10^8 =$$

Interpreting Graphics Ch 10 Chemistry Answers

L Cohen



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