

BETH "KEY"

Mole Conversions Worksheet

There are three mole equalities. They are:

1 mol = 6.022×10^{23} particles

1 mol = molar mass in grams (periodic table - red numbers)

1 mol = 22.4 L for a gas at STP

Mole-Particle Conversions

1. How many moles of magnesium are in 3.01×10^{22} atoms of magnesium?

$$3.01 \times 10^{22} \text{ atoms Mg} \times \frac{1 \text{ mol}}{6.022 \times 10^{23} \text{ atoms}} = 0.04998 \text{ mol Mg} \rightarrow \boxed{0.0500 \text{ mol Mg}}$$

2. How many molecules are there in 4.00 moles of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$?

$$4.00 \text{ mol C}_6\text{H}_{12}\text{O}_6 \times \frac{6.022 \times 10^{23} \text{ molecules}}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} = \boxed{2.41 \times 10^{24} \text{ molecules C}_6\text{H}_{12}\text{O}_6}$$

How many atoms of carbon?

$$2.41 \times 10^{24} \text{ molecules C}_6\text{H}_{12}\text{O}_6 \times \frac{6 \text{ atoms C}}{1 \text{ molecule C}_6\text{H}_{12}\text{O}_6} = \boxed{1.45 \times 10^{25} \text{ atoms C}}$$

How many atoms of hydrogen?

$$2.41 \times 10^{24} \text{ molecules C}_6\text{H}_{12}\text{O}_6 \times \frac{12 \text{ atoms H}}{1 \text{ molecule C}_6\text{H}_{12}\text{O}_6} = \boxed{2.89 \times 10^{25} \text{ atoms H}}$$

3. How many moles are 1.20×10^{25} atoms of phosphorous?

$$1.20 \times 10^{25} \text{ atoms P} \times \frac{1 \text{ mol P}}{6.022 \times 10^{23} \text{ atoms P}} = \boxed{19.9 \text{ mol P}}$$

4. How many atoms are in 0.750 moles of zinc?

$$0.750 \text{ mol Zn} \times \frac{6.022 \times 10^{23} \text{ atoms Zn}}{1 \text{ mol Zn}} = \boxed{4.52 \times 10^{23} \text{ atoms Zn}}$$

5. How many molecules are in 0.400 moles of N_2O_5 ?

$$0.400 \text{ mol N}_2\text{O}_5 \times \frac{6.022 \times 10^{23} \text{ molecules N}_2\text{O}_5}{1 \text{ mol N}_2\text{O}_5} = \boxed{2.41 \times 10^{23} \text{ molecules N}_2\text{O}_5}$$

How many atoms of nitrogen?

$$2.41 \times 10^{23} \text{ molecules N}_2\text{O}_5 \times \frac{2 \text{ atoms N}}{1 \text{ molecule N}_2\text{O}_5} = \boxed{4.82 \times 10^{23} \text{ atoms N}}$$

How many atoms of oxygen?

$$2.41 \times 10^{23} \text{ molecules N}_2\text{O}_5 \times \frac{5 \text{ atoms O}}{1 \text{ molecule N}_2\text{O}_5} = \boxed{1.21 \times 10^{24} \text{ atoms O}}$$

Moles Conversion Mole Map Answer Key

Bruce A. Finlayson

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Stephen DeMeo,2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this competitive environment entails not just studying harder or longer it means re organizing textbook content so that it is meaningful to the student This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career *Foundations of College Chemistry* Morris Hein,Susan Arena,Cary Willard,2023 *Foundations of College Chemistry* 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity *Chemistry* Paul B. Kelter,James D. Carr,Andrew Scott,1998 *Chemistry* Neil D. Jespersen,Alison Hyslop,2021-11-02 *Chemistry The Molecular Nature of Matter* 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they

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Chemistry Martin Stuart Silberberg,1996 This general chemistry text centres on the theme that observable change in chemical systems is the result of molecular change The aims of this edition are to enable students to perceive matter and change at the molecular level and to help build student confidence in their ability to solve chemical problems as they discover the relevance of chemistry to their lives Prentice Hall Chemistry ,2000 Doubly Labelled

Water J. Speakman,1997-09-30 Divided into three parts Doubly Labelled Water presents a clear and accessible account of this technique Part One presents a general introduction to the study of animal energetics Part Two discusses the theory behind use of doubled labelled water and Part Three evaluates the practical aspects of its use and the methodologies required for its application

Chemistry James C. Hill,2003 This book assists students through the text material with chapter overviews learning objectives review of key terms cumulative chapter review quizzes and self tests Included are answers to all Student Guide exercises Chapter summaries are correlated to those in the Instructor s Resource Manual

Cumulated Index Medicus ,1996

Thermodynamics Problem Solving in Physical Chemistry Kathleen E. Murphy,2020-03-23 Thermodynamics Problem Solving in Physical Chemistry Study Guide and Map is an innovative and unique workbook that guides physical chemistry students through the decision making process to assess a problem situation create appropriate solutions and gain confidence through practice solving physical chemistry problems The workbook includes six major sections with 20 30 solved problems in each section that span from easy single objective questions to difficult multistep analysis problems Each section of the workbook contains key points that highlight major features of the topic to remind students of what they need to apply to solve problems in the topic area Key Features Includes a visual map that shows how all the equations used in thermodynamics are connected and how they are derived from the three major energy laws Acts as a guide in deriving the correct solution to a problem Illustrates the questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry Can be used as a stand alone product for review of Thermodynamics questions for major tests

OCR(A) A2 Physics Student Unit Guide: Unit G485 Fields, Particles and Frontiers of Physics Gurinder Chadha,2012-01-01 Student Unit Guides are perfect for revision Each guide is written by an examiner and explains the unit requirements summarises the relevant unit content and includes a series of specimen questions and answers

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