

Molecule Shapes and Inter Molecular Forces

Introduction

Atoms bond to satisfy their need for more electrons. If both atoms have high electro-negativities (are nonmetals), atoms will share electrons to satisfy the Octet Rule – every atom wants 8 electrons to fill the s and p orbitals in the outer energy level. But, as you will see, if the electro-negativities are high enough and both atoms unwilling to give up electrons, sometimes atoms can deviate from and not follow) the Octet Rule.

Prelab:

Define following terms in your lab note book:

- | | |
|----------------------------|-----------------------|
| 1. VSEPR | 6. Electron geometry |
| 2. Electronegativity | 7. Molecular geometry |
| 3. Octet Rule | 8. Bond dipole |
| 4. Exception to octet rule | 9. Molecule polarity |
| 5. Electron domain | |

Molecule Shapes



Log on to <https://phet.colorado.edu/en/simulation/molecule-shapes> either by Googling "phet simulations molecule shape." Click on "**Run Now.**" If a screen pops up asking to update Java, click on "**Later.**"

Part 1: Electron domains

Explore the *Model* screen of the simulation and check what happens to existing atoms/lone pairs

- when you add an atom or lone pair
- when you remove an atom or lone pair,
- does adding an atom have same effect as adding a lone pair?

Part 2: 3D Drawings

- Line: In the plane of the paper: —
- Wedge: Coming forward, in front of the plane of the paper: 
- Dash: Going backward, behind the plane of the paper: 

- Draw the molecule of CHClBr on left in your notebook and label the position of each atom as

- In the plane of the paper _____
- In front of the plane of the paper _____
- Behind the plane of the paper _____



- Using the *Model* screen, add bonding domains (●) to the central atom (o). Using lines, wedges and dashes, draw each molecule's shape. Draw table in your lab notebook. Note: Bonding domains include any kind of bond and lone pairs.

Bonding Domains Around Central Atom	Drawing of Shape	Electron Geometry	Bond Angles
2	●—○—●	Linear	180°
3	○		

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LL Leslie

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