

SYLLABUS

Mole Concept and Stoichiometry

- (i) Gay-Lussac's Law of Combining Volumes; Avogadro's Law.
Idea of mole – a number just as dozen, a gross; Avogadro's Law – statement and explanation; Gay-Lussac's Law of Combining Volumes – statement and explanation; "the mass of 22.4 litres of any gas at S.T.P. is equal to its molar mass" (Questions will not be set on formal proof but may be taught for clear understanding) – simple calculations based on the molar volume.
- (ii) Refer to the atomicity of hydrogen, oxygen, nitrogen and chlorine (proof not required).
The explanation can be given using equations for the formation of HCl, NH₃ and NO.
- (iii) Relative atomic masses (atomic weight) and relative molecular masses (molecular weights): either H = 1 or ¹²C = 12 will be accepted; molecular mass = 2 × vapour density (formal proof not required). Deduction of simple (empirical) and molecular formula from the percentage composition of a compound; the molar volume of a gas at S.T.P.; simple calculations based on chemical equations; both reacting weight and volumes.
*Idea of relative atomic mass and relative molecular mass – standard H atom, or 1/12th of carbon 12 atom.
 Relating mole and atomic mass; arriving at gram atomic mass and then gram atom; atomic mass is a number dealing with one atom; gram atomic mass is the mass of one mole of atoms.
 Relating mole and molecular mass arriving at gram molecular mass and gram molecule – molecular mass is a number dealing with a molecule, gram molecular mass is the mass of one mole of molecules.
 Molecular mass = 2 × vapour density (questions will not be set on formal proof but may be taught for clear understanding); simple calculations based on the formula.
 Deduction of simple (empirical) and molecular formula from the percentage composition of a compound.*

5A. GAY-LUSSAC'S LAW AND AVOGADRO'S LAW

5.1 INTRODUCTION

You have learnt in Class IX that all gases behave similarly under similar conditions of temperature and pressure as expressed by gas laws.

Gas laws are certain rules applicable to a gas in respect of change in temperature (*T*) or pressure (*P*) or volume (*V*). The change in any one of these affects the other two.

Pressure-Volume Relationship or Boyle's Law

It states that *the volume of a given mass of dry gas is inversely proportional to its pressure at a constant temperature.*

$$P_1 V_1 = P_2 V_2 = k \text{ at constant temperature}$$

Temperature-Volume Relationship or Charles's Law :

It states that *volume of a given mass of a dry gas is directly proportional to its absolute (kelvin) temperature, if the pressure is kept constant.*

OR

The pressure remaining constant, the volume of a given mass of a dry gas increases or decreases by 1/273 of its volume for each 1°C increase or decrease in temperature respectively.

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} = k \text{ at constant pressure}$$

Gas Equation

On combining both the laws we conclude that

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R Bogdan



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Teaching the Mole Concept in Stoichiometry Timothy Scot Albrecht, California State Polytechnic University, Pomona. College of Education & Integrative Studies, 2009 Formulation and Stoichiometry Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction The Mole Concept Paul Gilchrist, 1982* **The Mole Concept**, 1983 SUMMARY A series of 6 programmes explaining the reasoning behind the development and acceptance of the mole concept **Stoichiometry** Sydney Young, 1908 **The Teaching of the Mole Concept** John Slater Allen, 1976 *The Mole Concept* Alan Keith Griffiths, 1979 **Students' Conceptions of Stoichiometry at the Submicro Level** Sin-Yan Chan, □□□, 2017-01-27 This dissertation Students Conceptions of Stoichiometry at the Submicro Level by Sin yan Chan was obtained from The University of Hong Kong Pokfulam Hong Kong and is being sold pursuant to Creative Commons Attribution 3 0 Hong Kong License The content of this dissertation has not been altered in any way We have altered the formatting in order to facilitate the ease of printing and reading of the dissertation All rights not granted by the above license are retained by the author Abstract Stoichiometry is an important topic in chemistry It tells how many reactants are required to produce a certain amount of product in terms of mass mole and volume Learning stoichiometric calculation involves the understanding of certain concepts such as the mole stoichiometric ratios and chemical equations Some studies attributed the failure in learning stoichiometry to the unfamiliarity with the amount of substance in a mole and students weakness in the mathematical ability Nevertheless how students connect the submicro level and the symbolic level in learning stoichiometry was not discussed widely In this study two examples of chemical reactions with different levels of difficulties were used to probe students conceptual understanding in stoichiometry at the submicro level Their strategies used in stoichiometric calculations were examined by an interview study of five Secondary Five students The connection between the submicro level and the symbolic level in

learning stoichiometry would be also probed into Results indicate that the failure of stoichiometry learning may due to the disconnection in different levels of representation and students generated strategy one portion reasoning An implication for teaching and learning is that teachers should use diagrams at the submicro level in the teaching of stoichiometry Such diagrams should aim to help students building connections across the three levels of representation and enhancing students conceptual understanding in stoichiometry DOI 10.5353/th_b5396387 Subjects Chemistry Study and teaching Secondary

The Mole Concept : Investigation of an Hierarchical Model Alan K. (Alan Keith) Griffiths, 1979

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Table of Contents Mole Concept And Stoichiometry Fy Bsc Chapter

1. Understanding the eBook Mole Concept And Stoichiometry Fy Bsc Chapter
 - The Rise of Digital Reading Mole Concept And Stoichiometry Fy Bsc Chapter
 - Advantages of eBooks Over Traditional Books
2. Identifying Mole Concept And Stoichiometry Fy Bsc Chapter
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mole Concept And Stoichiometry Fy Bsc Chapter
 - User-Friendly Interface

4. Exploring eBook Recommendations from Mole Concept And Stoichiometry Fy Bsc Chapter
 - Personalized Recommendations
 - Mole Concept And Stoichiometry Fy Bsc Chapter User Reviews and Ratings
 - Mole Concept And Stoichiometry Fy Bsc Chapter and Bestseller Lists
5. Accessing Mole Concept And Stoichiometry Fy Bsc Chapter Free and Paid eBooks
 - Mole Concept And Stoichiometry Fy Bsc Chapter Public Domain eBooks
 - Mole Concept And Stoichiometry Fy Bsc Chapter eBook Subscription Services
 - Mole Concept And Stoichiometry Fy Bsc Chapter Budget-Friendly Options
6. Navigating Mole Concept And Stoichiometry Fy Bsc Chapter eBook Formats
 - ePub, PDF, MOBI, and More
 - Mole Concept And Stoichiometry Fy Bsc Chapter Compatibility with Devices
 - Mole Concept And Stoichiometry Fy Bsc Chapter Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mole Concept And Stoichiometry Fy Bsc Chapter
 - Highlighting and Note-Taking Mole Concept And Stoichiometry Fy Bsc Chapter
 - Interactive Elements Mole Concept And Stoichiometry Fy Bsc Chapter
8. Staying Engaged with Mole Concept And Stoichiometry Fy Bsc Chapter
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole Concept And Stoichiometry Fy Bsc Chapter
9. Balancing eBooks and Physical Books Mole Concept And Stoichiometry Fy Bsc Chapter
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole Concept And Stoichiometry Fy Bsc Chapter
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mole Concept And Stoichiometry Fy Bsc Chapter
 - Setting Reading Goals Mole Concept And Stoichiometry Fy Bsc Chapter
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Mole Concept And Stoichiometry Fy Bsc Chapter
 - Fact-Checking eBook Content of Mole Concept And Stoichiometry Fy Bsc Chapter
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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