

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

Modern Chemistry Chapter 6 Chemical Bonding Test

Answer Key

Rob Ritchie



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Table of Contents Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key

1. Understanding the eBook Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - The Rise of Digital Reading Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - User-Friendly Interface

4. Exploring eBook Recommendations from Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Personalized Recommendations
 - Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key User Reviews and Ratings
 - Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key and Bestseller Lists
5. Accessing Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key Free and Paid eBooks
 - Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key Public Domain eBooks
 - Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key eBook Subscription Services
 - Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key Budget-Friendly Options
6. Navigating Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key Compatibility with Devices
 - Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Highlighting and Note-Taking Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Interactive Elements Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
8. Staying Engaged with Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
9. Balancing eBooks and Physical Books Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Setting Reading Goals Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Fact-Checking eBook Content of Modern Chemistry Chapter 6 Chemical Bonding Test Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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