

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

a. Sulfur hexafluoride  $SF_6$

b. Lithium nitride  $Li_3N$

c. Chromium (III) carbonate  $Cr_2(CO_3)_3$

d. Tin (II) chloride  $TnCl_2$

e. Ammonium acetate  $NH_4C_2H_3O_2$  (or  $CH_3COONH_4$ )

f. Mercury (I) chloride  $Hg_2Cl_2$

g. Potassium bromate  $KBrO_3$

h. Hydrosulfuric acid  $H_2S(aq)$

i. Chloric acid  $HClO_3$

j. Sulfurous acid  $H_2SO_3$

Name each of the following compounds.

a. CuSO<sub>4</sub> Copper (II) sulfate

b. AlF<sub>3</sub> aluminum fluoride

c. HI hydrogen iodide

d. NO nitrogen monoxide

e. H<sub>2</sub>Se(aq) hydroseconic acid

f. HNO<sub>3</sub> nitric acid

g. HNO<sub>2</sub> nitrous acid

h. NaHSO<sub>4</sub> sodium bisulfate

### Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it takes for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives } 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20 \\ & + & \\ & & 63.62 \rightarrow 64 \end{array}$$

# Modern Chemistry Answer Key 8

**Dr. Satkar Prasad, Dr. Ganesh S  
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