

AQA Chemistry

GCSE Exam-style mark scheme

C4 Chemical calculations

100 marks (20%)

1 hour

Question	Answer	Mark	Mark	Mark
10.1.1	1.4 g		1	100
10.1.2	1.48 g	Amount of magnesium carbonate $1.48 \text{ g} \div 142.04 \text{ g mol}^{-1}$ $= 0.01042 \text{ mol}$ $\times 1000 = 10.42 \text{ mmol}$ <u>10</u> mol as amount of Ca^{2+} ions bromine ions should be produced by 2 moles	10	100
10.1.3	very high temperature releasing energy used as magnesium carbonate used as magnesium oxide from the reaction		10	100
10.2	all compounds used are changed amounts all used 10 g so the same amount of all will contribute same mass of oxygen atoms in water (oxygen added hydrogen)	Mass of hydrogen in water hydrogen given 10 g of hydrogen in water mass of hydrogen in water is 10 g because the water is 10 g	10 10 10	100
10.3.1	the water molecules are being transferred to water		10 10	100
10.3.2	1.4 g of 1.48 g 1.4 g of 1.48 g of 1.48 g 1.4 g of 1.48 g 1.4 g of 1.48 g	Mass 1.4 g of 1.48 g of 1.48 g 1.4 g of 1.48 g Mass 1.4 g of 1.48 g of 1.48 g 1.4 g of 1.48 g of 1.48 g	10 10 10	100

Mei June 2013 C4 Mark Scheme

Yijin Wang



Mei June 2013 C4 Mark Scheme:

Mei June 2013 C4 Mark Scheme Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Mei June 2013 C4 Mark Scheme**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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