

Study Guide for Honors Chemistry Final - June 2015

You will be given all needed equations, values of constants and other pertinent information, including the following: periodic table, $c = \lambda\nu$, $E = h\nu$, the value for $h = 6.626 \times 10^{-34}$ J·s, the conversion between meters and nanometers ($1 \text{ m} = 1 \times 10^9 \text{ nm}$), electronegativity values, activity series of metals, gas law equations, dilution equation, molarity, molality. It is your responsibility to recognize the equations and understand how to use them. You will also be given a 3" x 5" notecard that, must be handwritten by yourself, you may use to organize your thoughts for the exam.

Do the sample problems and review all the vocabulary words in the sections indicated. Go to the following links to test your knowledge on the Chapters (remember we did not do all of the sections in all of the following chapters):

http://chemistrymc.com/chapter_test and http://chemistrymc.com/standardized_test

Electrons in Atoms – Chapter 5

Objectives:

- Describe the relationship between wavelength and frequency.
- Be able to use the equation $c = \lambda\nu$ where $c = 3.00 \times 10^8$ m/s
- Describe the relationship between frequency and energy of a wave.
- Be able to use the equation $E = h\nu$ where $h = \text{Planck's constant} = 6.626 \times 10^{-34}$ J·s
- Describe the contributions of Bohr, Schrödinger, Planck, Einstein, & deBroglie to determining the location of electrons in the atom and their properties
- Explain the difference between a continuous spectrum and a line spectrum and give examples of each.
- Be able to arrange the colors of the visible spectrum from lowest energy to highest energy.
- Define a quantum of energy and explain how it is related to an energy change of matter.
- Describe how the quantum mechanical model describes what an atom is and where the subatomic particles are located.
- Explain the relationships among an atom's energy level, sublevels, and atomic orbitals.
- Apply the Pauli Exclusion Principle, the Aufbau Principle, and Hund's Rule to write electron configurations using orbital diagrams and electron configuration notation.
- Define valence electrons and draw electron-dot structures representing an atom's valence electrons.
- Be able to read and write quantum numbers given an electron configuration or an element's symbol
- Labs – Thickness of Zinc, Spectroscopy and Flame Tests, Electron Arrangements

Vocabulary:

- | | |
|---|--|
| • Wavelength, frequency, amplitude and speed of waves | • Principal energy level |
| • Electromagnetic spectrum | • Orbitals (s, p, d, and f) |
| • Quantum | • Electron configuration |
| • Photoelectric effect | • Valence electron |
| • Atomic emission spectrum | • Electron-dot structures (Lewis dot diagrams) |
| • Ground state vs. excited state | • Quantum numbers |
| • Heisenberg uncertainty principle | • $n, \ell, m\ell, m_s$ |
| • Quantum mechanical model of the atom | |
| • Atomic orbital | |
| • Principal quantum number | |

Honors Chemistry Final Study Guide

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Honors Chemistry Final Study Guide:

The Enigmatic Realm of **Honors Chemistry Final Study Guide**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Honors Chemistry Final Study Guide** a literary masterpiece penned with a renowned author, readers embark on a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting impact on the hearts and minds of people who partake in its reading experience.

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