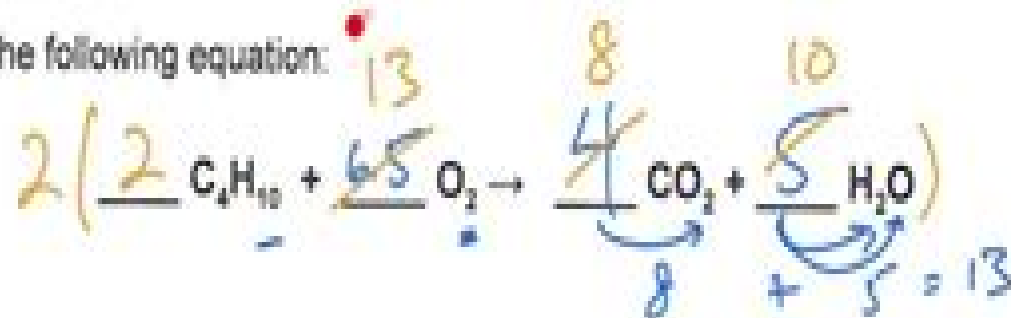


Balance the following equation:



1. How many moles are in  $4.52 \times 10^{30}$  molecules of  $\text{O}_2$ ?

$$4.52 \times 10^{30} \text{ molecules} \times \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} = 7.51 \times 10^6 \text{ mol O}_2$$

2. If you have 3.5 moles of  $\text{O}_2$ , how many moles of  $\text{H}_2\text{O}$  will you make?

$$3.5 \text{ mol O}_2 \times \frac{10 \text{ mol H}_2\text{O}}{13 \text{ mol O}_2} = \text{mol H}_2\text{O}$$

3. If you have 50.0 g of  $\text{C}_4\text{H}_{10}$ , how many moles of  $\text{O}_2$  will you need?

4. If you have 100.0 g of  $\text{O}_2$ , how many grams of  $\text{CO}_2$  will you make?

# Honor Chem Review Exam

**S Ashworth**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the name S Ashworth.

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