

DT92E Wireless Room Thermostat with ECO energy saving feature

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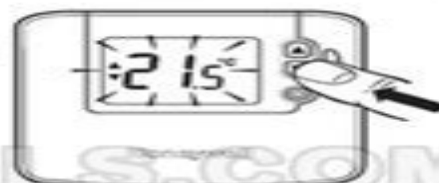
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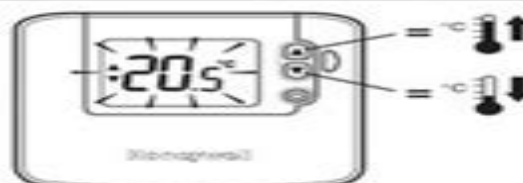
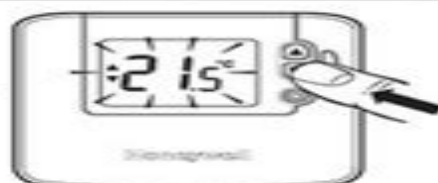


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F Lecture de la température demandée
NL Uitlezen van de ingestelde temperatuur
D Eingestellte Temperatur anzeigen
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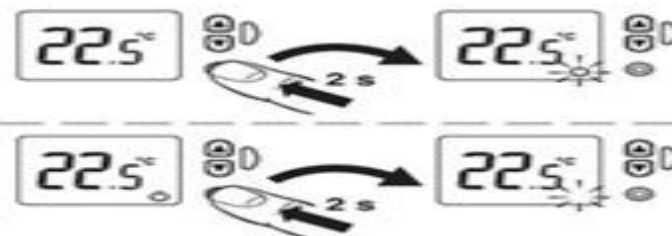
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PL Odczyt temperatury zadanej
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- GB Switch ON/OFF (frost protection enabled)
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F Bouton Marche/Arrêt (avec protection anti-gel)
NL AAN/UIT-toets (vorstbeveiliging actief)
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web 1 78 348 stoichiometry helps us determine how much substance is needed or is present things that can be measured are reactants and products mass molecular weight chemical equations formulas stoichiometric coefficient the stoichiometric coefficient or stoichiometric number is the number of molecules that participate in the reaction

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n2 f2 nfs 2 c6h10 02 co2 h2o 3 hbr khco3 h2o kbr co2 gabr nabr sno 4 naso3 gaz so3 3 5 nf3 snf2 n2o3 using the equation
from problem 2 above answer the following

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web balancing equations and simple stoichiometry key balance the following equations

$$\begin{aligned} & \text{1 C}_6\text{H}_{10}\text{O}_2 + \text{12 CO}_2 + \text{10 H}_2\text{O} + \text{3 O}_2 \rightarrow \text{2 NF}_3 + \text{2 C}_6\text{H}_6 + \text{17 O}_2 + \text{12 CO}_2 \\ & \text{10 H}_2\text{O} + \text{3 O}_2 + \text{5 HBr} + \text{1 KHCO}_3 + \text{1 H}_2\text{O} + \text{1 KBr} + \text{2 CO}_2 + \text{2 GaBr}_3 + \text{3 Na}_2\text{SO}_3 + \text{1 Ga}_2\text{SO}_3 + \text{3 H}_2\text{O} + \text{6 NaBr} + \text{3 SnO}_2 + \text{3 NF}_3 + \text{1 SnF}_2 + \text{1 N}_2\text{O}_3 \end{aligned}$$

using the following equation

$$\text{2 NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{2 H}_2\text{O} + \text{Na}_2\text{SO}_4$$

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assuming CeCl_2 CuCl_2 is available in excess

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 $\text{Br}_2 + \text{C}_6\text{H}_{10} \rightarrow \text{C}_6\text{H}_8 + \text{HBr}$
 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{Na}_2\text{SO}_4$
br using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
 $\text{CaH}_2 + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$
 $\text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{Na}_2\text{SO}_3$
 $\text{NaBr} + \text{NF}_3 \rightarrow \text{NF}_2$

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web may 28 2020 write the balanced equation for the reaction and calculate how much SiO_2 is required to produce 3.00 kg of SiC
 $\text{SiO}_2 + \text{C} \rightarrow \text{SiC} + \text{CO}$

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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
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 $\text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{Na}_2\text{SO}_3$
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same thing goes with the o3 i am so lost

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