

History of the Atom



Democritus

He proposed that all matter, including space and time, was made up of small units named atoms. He did no experiments and had little evidence, but his idea was kept on by Lucretius



Mendeleev

Created the periodic table of elements which organizes elements according to their similarities. His Period Law states that "physical and chemical of the elements are periodic functions of their atomic numbers."



Henri Becquerel

Discovered radioactivity, which earned him a Nobel Prize. Also discovered that rays emitted from uranium caused gases to ionize



Ernest Rutherford

Used the gold foil experiment to discover the modern model of an atom. He concluded that all positive charges were centralized, while negative electrons orbited the nucleus



Niels Bohr

Explained that outer orbits in an atom could hold more electrons than the inner orbits. By knowing this, one can determine the atom's chemical properties. He also gave birth to the idea that electrons emit light by jumping orbits.



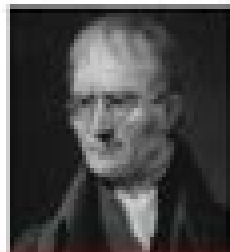
Heisenberg

Thought of the Uncertainty Principle. This principle states that one can never know the exact location and energy of an electron simultaneously.



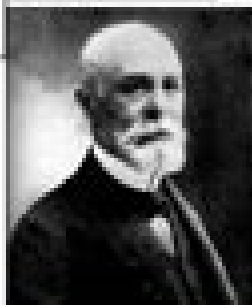
Dalton

Agreed that all matter was created by atoms, which he believed were indestructible. He also stated that compounds are created by combining two atoms, and that all atoms of given elements are identical in their mass and properties.



Eugene Goldstein

Discovered positive particles. He noted that the particles had a charge equal and opposite to the electron.



J.J. Thomson

Discovered the electron. He experimented by testing and studying the nature of electric discharge in a high vacuum cathode-ray tube.



Millikan

Used an oil drop experiment to determine the charge of a single electron.



Frederick Soddy

Came up with the concept of isotopes. Explain with Ernest Rutherford that radiation is due to transmutation of elements.



Chadwick

Discovered the neutron. The neutron helps balance out protons in the nucleus of an atom.



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Wolfgang Guggemos



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Table of Contents History Of The Atom Notes Key Tesccc

1. Understanding the eBook History Of The Atom Notes Key Tesccc
 - The Rise of Digital Reading History Of The Atom Notes Key Tesccc
 - Advantages of eBooks Over Traditional Books
2. Identifying History Of The Atom Notes Key Tesccc
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an History Of The Atom Notes Key Tesccc
 - User-Friendly Interface
4. Exploring eBook Recommendations from History Of The Atom Notes Key Tesccc
 - Personalized Recommendations
 - History Of The Atom Notes Key Tesccc User Reviews and Ratings
 - History Of The Atom Notes Key Tesccc and Bestseller Lists

5. Accessing History Of The Atom Notes Key Tesccc Free and Paid eBooks
 - History Of The Atom Notes Key Tesccc Public Domain eBooks
 - History Of The Atom Notes Key Tesccc eBook Subscription Services
 - History Of The Atom Notes Key Tesccc Budget-Friendly Options
6. Navigating History Of The Atom Notes Key Tesccc eBook Formats
 - ePub, PDF, MOBI, and More
 - History Of The Atom Notes Key Tesccc Compatibility with Devices
 - History Of The Atom Notes Key Tesccc Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of History Of The Atom Notes Key Tesccc
 - Highlighting and Note-Taking History Of The Atom Notes Key Tesccc
 - Interactive Elements History Of The Atom Notes Key Tesccc
8. Staying Engaged with History Of The Atom Notes Key Tesccc
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers History Of The Atom Notes Key Tesccc
9. Balancing eBooks and Physical Books History Of The Atom Notes Key Tesccc
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection History Of The Atom Notes Key Tesccc
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine History Of The Atom Notes Key Tesccc
 - Setting Reading Goals History Of The Atom Notes Key Tesccc
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of History Of The Atom Notes Key Tesccc
 - Fact-Checking eBook Content of History Of The Atom Notes Key Tesccc
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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