



NATIONAL  
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2  
(CHEMISTRY)  
EXEMPLAR

Duration: 120  
Time: 1 hour



The question paper consists of 12 pages, including 4 extra sheets.  
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## QUESTION 8

### PHYSICAL SCIENCES PAPER 2

### ACID AND BASES

**GRADE 11**

**NOVEMBER 2020**

QUESTION 8: ACID AND BASES (15)

8.1. Give a balanced ionic equation for part (b) of the reaction in question 6.

The following species are shown on the page:

- a.  $\text{H}^+$  (aq) (1)  
b.  $\text{H}_2\text{O}$  (l) (1)

QUESTION 8 (Start on a NEW page)

8.2. Nitric acid reacts with an organic compound with the following percentage composition:

C: 55.4%, H: 4.54%, N: 39.96%

The molar mass of the compound is 74 g mol<sup>-1</sup>.

- 8.2.1. Deduce the full empirical formula. (3)  
8.2.2. Give the IUPAC name of the compound. (1)  
8.2.3. Give the molecular formula. (1)  
8.3. Concentrated nitric acid reacts with sodium hydroxide solution to form sodium nitrate. 2 g of sodium hydroxide (NaOH) is a substance that can release water in a process called a neutralisation reaction.  
8.3.1. Deduce the concentration of the sodium hydroxide solution. (3)  
8.3.2. Calculate the concentration of the sodium hydroxide solution. (3)  
8.4. Sodium nitrate (NaNO<sub>3</sub>) is used in car airbags. For the nitrate ion, write the following reaction with the acid.

$\text{H}^+$  (aq) +  $\text{NO}_3^-$  (aq) →

Calculate the volume of nitrogen gas (N<sub>2</sub>) that would be produced at STP if 10 g of sodium nitrate reacts completely. (3)



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**Christian Drosten**



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