

PHYSICAL SCIENCES – GRADE 10

CHAPTER – 1 Chemical Reactions and Equations

- q **Chemical Reaction** : – Whenever a chemical change occurs we can say that a chemical reaction has taken place
eg – Food gets digested in our body
 – Rusting of iron.
- q **Chemical Equation** :– A chemical reaction can be expressed symbolically by using chemical equation
eg magnesium is burnt into air to form magnesium oxide can be represented as
$$\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$$

– We can observe or recognise a chemical reaction by observing change in state, colour, by evolution of gas or by change in temperature.
- q Physical state of the reactant and products are mentioned to make chemical reaction more informative. eg we use (g) for gas, (l) for liquid, (s) for solid and (aq) for aqueous.
- q **Balancing Equation** :– We balance the chemical equation so that no. of atoms of each element involved in the reaction remain same at the reactant and product side.
eg $\text{Fe} + \text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + \text{H}_2$ can be written as
$$3 \text{Fe(s)} + 4\text{H}_2\text{O(g)} \rightarrow \text{Fe}_3\text{O}_4\text{(s)} + 4\text{H}_2\text{(g)}$$
- q **Combination Reaction** :– The reaction in which two or more substances combine to form a new single substance

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1. Understanding the eBook Grade10 Physical Science P2 June Exam 2015
 - The Rise of Digital Reading Grade10 Physical Science P2 June Exam 2015
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade10 Physical Science P2 June Exam 2015
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade10 Physical Science P2 June Exam 2015
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade10 Physical Science P2 June Exam 2015
 - Personalized Recommendations
 - Grade10 Physical Science P2 June Exam 2015 User Reviews and Ratings
 - Grade10 Physical Science P2 June Exam 2015 and Bestseller Lists
5. Accessing Grade10 Physical Science P2 June Exam 2015 Free and Paid eBooks
 - Grade10 Physical Science P2 June Exam 2015 Public Domain eBooks
 - Grade10 Physical Science P2 June Exam 2015 eBook Subscription Services
 - Grade10 Physical Science P2 June Exam 2015 Budget-Friendly Options
6. Navigating Grade10 Physical Science P2 June Exam 2015 eBook Formats

- ePub, PDF, MOBI, and More
- Grade10 Physical Science P2 June Exam 2015 Compatibility with Devices
- Grade10 Physical Science P2 June Exam 2015 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade10 Physical Science P2 June Exam 2015
 - Highlighting and Note-Taking Grade10 Physical Science P2 June Exam 2015
 - Interactive Elements Grade10 Physical Science P2 June Exam 2015
- 8. Staying Engaged with Grade10 Physical Science P2 June Exam 2015
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade10 Physical Science P2 June Exam 2015
- 9. Balancing eBooks and Physical Books Grade10 Physical Science P2 June Exam 2015
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade10 Physical Science P2 June Exam 2015
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Grade10 Physical Science P2 June Exam 2015
 - Setting Reading Goals Grade10 Physical Science P2 June Exam 2015
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade10 Physical Science P2 June Exam 2015
 - Fact-Checking eBook Content of Grade10 Physical Science P2 June Exam 2015
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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