

Chemistry – Unit 6 Worksheet 1

We have observed evidence that atoms of different elements are different, the most noticeable being in their naturally-occurring isotopes. Atoms of the same element, which would be expected to have identical chemical properties, have different masses. This is because isotopes of an element have different numbers of neutrons in their nuclei. We will now examine the patterns that exist for the ratios of atomic masses of different elements in order to determine the charges of the ions they form.

1. Write the formula and draw the particle diagram for each compound. The ratio of ions in each compound is given.

Atomic number of element	16 electrons 16 protons $16/16 = 1$	16 electrons 14 protons $16/14 = 1.14$	16 electrons 12 protons $16/12 = 1.33$	16 electrons 10 protons $16/10 = 1.6$
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particle



diagram

Atomic number of element	16 electrons 14 protons $16/14 = 1.14$	16 electrons 12 protons $16/12 = 1.33$	16 electrons 10 protons $16/10 = 1.6$
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particle



diagram

Atomic number of element	16 negatively-charged 16 electron ions $16/16 = 1$	16 electrons 14 electron ions $16/14 = 1.14$	16 negatively-charged 12 electron ions $16/12 = 1.33$	16 electrons 10 electron ions $16/10 = 1.6$
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particle



diagram

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Yan Bai



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