

BETH "KEY"

Mole Conversions Worksheet

There are three mole equalities. They are:

1 mol = 6.022×10^{23} particles

1 mol = molar mass in grams (periodic table - red numbers)

1 mol = 22.4 L for a gas at STP

Mole-Particle Conversions

1. How many moles of magnesium are in 3.01×10^{22} atoms of magnesium?

$$3.01 \times 10^{22} \text{ atoms Mg} \times \frac{1 \text{ mol}}{6.022 \times 10^{23} \text{ atoms Mg}} = 0.04998 \text{ mol Mg} \rightarrow \boxed{0.0500 \text{ mol Mg}}$$

2. How many molecules are there in 4.00 moles of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$?

$$4.00 \text{ mol C}_6\text{H}_{12}\text{O}_6 \times \frac{6.022 \times 10^{23} \text{ molecules}}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} = \boxed{2.41 \times 10^{24} \text{ molecules C}_6\text{H}_{12}\text{O}_6}$$

How many atoms of carbon?

$$2.41 \times 10^{24} \text{ molecules C}_6\text{H}_{12}\text{O}_6 \times \frac{6 \text{ atoms C}}{1 \text{ molecule C}_6\text{H}_{12}\text{O}_6} = \boxed{1.45 \times 10^{25} \text{ atoms C}}$$

How many atoms of hydrogen?

$$2.41 \times 10^{24} \text{ molecules C}_6\text{H}_{12}\text{O}_6 \times \frac{12 \text{ atoms H}}{1 \text{ molecule C}_6\text{H}_{12}\text{O}_6} = \boxed{2.89 \times 10^{25} \text{ atoms H}}$$

3. How many moles are 1.20×10^{25} atoms of phosphorous?

$$1.20 \times 10^{25} \text{ atoms P} \times \frac{1 \text{ mol P}}{6.022 \times 10^{23} \text{ atoms P}} = \boxed{19.9 \text{ mol P}}$$

4. How many atoms are in 0.750 moles of zinc?

$$0.750 \text{ mol Zn} \times \frac{6.022 \times 10^{23} \text{ atoms Zn}}{1 \text{ mol Zn}} = \boxed{4.52 \times 10^{23} \text{ atoms Zn}}$$

5. How many molecules are in 0.400 moles of N_2O_5 ?

$$0.400 \text{ mol N}_2\text{O}_5 \times \frac{6.022 \times 10^{23} \text{ molecules N}_2\text{O}_5}{1 \text{ mol N}_2\text{O}_5} = \boxed{2.41 \times 10^{23} \text{ molecules N}_2\text{O}_5}$$

How many atoms of nitrogen?

$$2.41 \times 10^{23} \text{ molecules N}_2\text{O}_5 \times \frac{2 \text{ atoms N}}{1 \text{ molecule N}_2\text{O}_5} = \boxed{4.82 \times 10^{23} \text{ atoms N}}$$

How many atoms of oxygen?

$$2.41 \times 10^{23} \text{ molecules N}_2\text{O}_5 \times \frac{5 \text{ atoms O}}{1 \text{ molecule N}_2\text{O}_5} = \boxed{1.21 \times 10^{24} \text{ atoms O}}$$

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Neil D. Jespersen, Alison Hyslop



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