

Moles and Chemical Formulas

11

A. Finding the Simplest Formula

1. Mass of empty crucible + cover
2. Initial appearance of the magnesium
3. Mass of crucible + cover + magnesium
4. Mass of crucible + cover + oxide product

	<u>22.842</u>	g
	<u>23.483</u>	
	<u>23.073</u>	g
	<u>29.317</u>	g

Calculations

5. Mass of magnesium
6. Mass of magnesium compound
7. Mass of oxygen in the product
8. Moles of Mg

	<u>0.2305</u>	g
	<u>6.475</u>	g
	<u>6.245</u>	g
	<u>.0095</u>	g

(Show calculations.)

$$0.2305 \text{ g/mg} \times \frac{1 \text{ mole mg}}{24.31 \text{ g/mg}} = 0.009481$$

9. Moles of O

(Show calculations.)

$$6.245 \text{ g} \times \frac{1 \text{ mole}}{16.00} = 0.3903$$

10. Which number of moles (Mg or O) is smaller Mole of Mg
moles of Mg (rounded)

Moles And Chemical Formulas Lab Answers

**Marcia L. Gillette,Ira Blei,Wendy
Gloffke,George Odian**



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