

## Unit 7: Notes - Introducing Empirical and Molecular Formulas

Molecular Formula is a formula indicating the actual number of atoms of each element making up a molecule. **CAN BE REDUCED**!!!

Empirical formula is the formula giving the simplest ratio between the atoms of the elements present in a compound. **CAN NOT BE REDUCED**

A. Identify the following compounds as Empirical or Molecular

|                   |       |                         |       |
|-------------------|-------|-------------------------|-------|
| 1) $P_4O_6$       | _____ | 5) $C_6H_8O_6$          | _____ |
| 2) $C_6H_9$       | _____ | 6) $CuC_2O_4$           | _____ |
| 3) $CH_2OHCH_2OH$ | _____ | 7) $Hg_2F_2$            | _____ |
| 4) $BrCl_2$       | _____ | 8) $C_{22}H_{22}O_{12}$ | _____ |

### Converting Molecular Formula to Empirical Formula

Example 1: Hydrogen Peroxide; Molecular Formula is  $H_2O_2$  → subscripts are divisible by 2  
Empirical Formula is  $HO$

Example 2: Glucose; Molecular Formula is  $C_6H_{12}O_6$  → subscripts are divisible by 6  
Empirical Formula is  $CH_2O$

B. Reduce the Molecular Formulas below to their simplest Empirical Formula

- 9)  $P_4O_6$  \_\_\_\_\_
- 10)  $C_6H_9$  \_\_\_\_\_
- 11)  $C_6H_8O_6$  \_\_\_\_\_
- 12)  $Hg_2F_2$  \_\_\_\_\_
- 13)  $C_4H_8O_2$  \_\_\_\_\_
- 14)  $C_4H_4O_4$  \_\_\_\_\_

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**LM Reder**



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