

Key.

b) 6.45 moles Cr_2O_3 $\text{FW} = 152.0 \text{ g/mol}$ 980 g

$$6.45 \text{ mol} \times \left(\frac{152.0 \text{ g}}{1 \text{ mol}} \right) = 980 \text{ g}$$

c) 0.144 mole K_2S $\text{FW} = 110.3 \text{ g/mol}$ 15.9 g

$$0.144 \text{ mol} \times \left(\frac{110.3 \text{ g}}{1 \text{ mol}} \right) = 15.9 \text{ g}$$

d) 5.0 moles $\text{C}_6\text{H}_{12}\text{O}_6$ 180.0 g/mol 900 g $\approx 9.0 \times 10^2 \text{ g}$

$$5.0 \text{ mol} \times \left(\frac{180.0 \text{ g}}{1 \text{ mol}} \right) = 900 \text{ g}$$

e) 5.10×10^{-4} mole NaCl $\text{FW} = 58.5 \text{ g/mol}$ $0.0298 \text{ g} \approx 2.98 \times 10^{-2} \text{ g}$

$$5.10 \times 10^{-4} \text{ mol} \times \left(\frac{58.5 \text{ g}}{1 \text{ mol}} \right) = 0.0298 \text{ g}$$

Solve the following stoichiometry problems. Write a balanced equation for each.

a) How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas, to produce hydrogen chloride gas? Equation: $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow 2 \text{HCl}(\text{g})$

$$15.1 \text{ g} \times \left(\frac{1 \text{ mol}}{71.0 \text{ g}} \right) = 0.213 \text{ mol } \text{Cl}_2$$

$$0.213 \text{ mol } \text{Cl}_2 \left(\frac{1 \text{ mol H}_2}{1 \text{ mol Cl}_2} \right) = 0.213 \text{ mol H}_2 \quad \underline{0.213 \text{ mol H}_2}$$

b) How many grams of oxygen gas are needed to react completely with 16.2 g of hydrogen gas, to produce water? Equation: $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{l})$

$$16.2 \text{ g} \times \left(\frac{1 \text{ mol}}{2.0 \text{ g}} \right) = 8.10 \text{ mol H}_2$$

$$8.10 \text{ mol H}_2 \left(\frac{1 \text{ mol O}_2}{2 \text{ mol H}_2} \right) = 4.05 \text{ mol O}_2 \quad \underline{130 \text{ g O}_2}$$

$$4.05 \text{ mol O}_2 \left(\frac{32.0 \text{ g/mol}}{1 \text{ mol}} \right) = 130 \text{ g O}_2$$

c) How many moles of calcium oxide are produced when 36.5 g of calcium react completely with oxygen gas? Equation: $2 \text{Ca}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2 \text{CaO}(\text{s})$

$$36.5 \text{ g} \left(\frac{1 \text{ mol}}{40.1 \text{ g}} \right) = 0.910 \text{ mol Ca}$$

$$0.910 \text{ mol Ca} \left(\frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} \right) = 0.910 \text{ mol } \underline{0.910 \text{ mol CaO}}$$

d) How many grams of aluminum are required to produce 410 g of aluminum oxide, through reaction with oxygen gas? Equation: $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$

$$410 \text{ g} \left(\frac{1 \text{ mol}}{102.0 \text{ g}} \right) = 4.02 \text{ mol Al}_2\text{O}_3$$

$$4.02 \text{ mol Al}_2\text{O}_3 \left(\frac{4 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \right) = 8.04 \text{ mol Al}$$

$$8.04 \text{ mol Al} \left(\frac{27.0 \text{ g}}{1 \text{ mol}} \right) = 217 \text{ g}$$

$$\underline{217 \text{ g} \approx 220 \text{ g (2sf)}}$$

e) How many grams of solid iodine are produced by the decomposition of 25.4 grams of hydrogen iodide gas? Equation: $2 \text{HI} \rightarrow \text{H}_2 + \text{I}_2$

$$25.4 \text{ g} \left(\frac{1 \text{ mol}}{127.9 \text{ g}} \right) = 0.199 \text{ mol HI}$$

$$0.199 \text{ mol HI} \left(\frac{1 \text{ mol I}_2}{2 \text{ mol HI}} \right) = 0.0993 \text{ mol I}_2 \quad \underline{25.2 \text{ g I}_2}$$

$$0.0993 \text{ mol I}_2 \left(\frac{253.8 \text{ g}}{1 \text{ mol}} \right) = 25.2 \text{ g I}_2$$

Mole Ratio Wkst Answers

Arshad Iqbal

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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Table of Contents Mole Ratio Wkst Answers

1. Understanding the eBook Mole Ratio Wkst Answers
 - The Rise of Digital Reading Mole Ratio Wkst Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mole Ratio Wkst Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mole Ratio Wkst Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Ratio Wkst Answers
 - Personalized Recommendations
 - Mole Ratio Wkst Answers User Reviews and Ratings

- Mole Ratio Wkst Answers and Bestseller Lists
- 5. Accessing Mole Ratio Wkst Answers Free and Paid eBooks
 - Mole Ratio Wkst Answers Public Domain eBooks
 - Mole Ratio Wkst Answers eBook Subscription Services
 - Mole Ratio Wkst Answers Budget-Friendly Options
- 6. Navigating Mole Ratio Wkst Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mole Ratio Wkst Answers Compatibility with Devices
 - Mole Ratio Wkst Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mole Ratio Wkst Answers
 - Highlighting and Note-Taking Mole Ratio Wkst Answers
 - Interactive Elements Mole Ratio Wkst Answers
- 8. Staying Engaged with Mole Ratio Wkst Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole Ratio Wkst Answers
- 9. Balancing eBooks and Physical Books Mole Ratio Wkst Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole Ratio Wkst Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Ratio Wkst Answers
 - Setting Reading Goals Mole Ratio Wkst Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Ratio Wkst Answers
 - Fact-Checking eBook Content of Mole Ratio Wkst Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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