Name: _____ Class: _____ Date: _____

Molar Mass and Mole Calculations Worksheet

1. Calculate the molar mass for each of the following compounds:

a. $NaCl$ = 58.44 g/mol

b. Fe_2O_3 = 159.7 g/mol

c. $MgCO_3$ = 84.32 g/mol

d. $Ba(OH)_2$ = 171.35 g/mol

e. $(NH_4)_3PO_4$ = 149.12 g/mol

f. $Ca_3(PO_4)_2$ =310.18 g/mol

g. $C_6H_{12}O_6$ = 180.18 g/mol

h. C_5H_5N = 79.11 g/mol

2. Calculate the number of moles in each of the following masses:

a. 45.0 g of acetic acid, CH_3COOH = 0.75 mol

b. 7.04 g of lead (II) nitrate = 0.02 mol

3. Calculate the mass in grams of each of the following amounts:

a. 3.00 mol of selenium oxybromide, $SeOBr_2$ = 764.31 g
M=78.97+ 16+2(79.9)= 254.77 g/mol 254.77 x 3=

b. 488 mol of calcium carbonate = 48,843.92 g
M=40.08+12.01+3(16)= 100.09 g/mol 100.09 x 488=

4. Calculate the number of molecules or formula units in each of the following amounts:

a. 4.27 mol of tungsten(VI) oxide, WO_3 = 2.57×10^{24}
$4.27 \times 6.02 \times 10^{23}$=

b. 0.989 mol of nitrobenzene $C_6H_5NO_2$ = 5.95×10^{23}
$0.989 \times 6.02 \times 10^{23}$=

5. Calculate the mass of each of the following quantities:

a. 8.39×10^{23} molecules of fluorine = 52.96 g
$8.39 \times 19^{23}/6.02 \times 10^{23}$= 1.39 mol 1.39 x 38=
M= 2(19) = 38 g/mol

b. 6.82×10^{24} formula units of beryllium sulfate = 1,190.33 g
M= 9.01+ 32.06+ 4(16) = 105.07 105.07 x 11.33=
$6.82 \times 10^{24}/6.02 \times 10^{23}$= 11.33

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