

# Dilutions and Molarity

**Dilution** is the process used to make a new solution that is **less concentrated** than the original solution by adding more **solvent**.

$M_1$  = Initial Molarity

$V_1$  = Initial Volume

$$M_1 V_1 = M_2 V_2$$

$M_2$  = Final Molarity

$V_2$  = Final Volume

## Examples Questions:

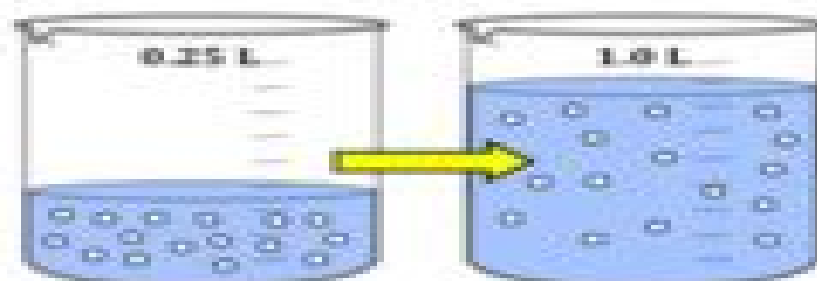
- 1) A scientist obtains 0.25 liters of concentrated 5.0 M NaCl solution. The scientist then **dilutes** the solution by adding water until the solution has a volume of 1.0 liters. What is the **molarity** of the solution?

$$M_1 V_1 = M_2 V_2$$

$$(5.0) (0.25) = M_2 (1.0)$$

$$(1.25) = M_2 (1.0)$$

$$M_2 = 1.25 \text{ M or mol/L}$$



- 2) A concentrated solution of 10.0 M hydrochloric acid solution must be **diluted** to make it weaker. If a scientist starts with 1.5 liters of the original solution and adds 2.5 liters of water, what will be the final **molarity** of the solution?

$$M_1 V_1 = M_2 V_2$$

$$(10.0) (1.5) = M_2 (4.0)$$

$$(15.0) = M_2 (4.0)$$

$$M_2 = 3.75 \text{ M or mol/L}$$

- 3) In a laboratory, a scientist takes 50 mL of 5.0 M NaOH solution and adds water until the solution has a molarity of 2.0 M. What is the **volume** of the new solution?

$$M_1 V_1 = M_2 V_2$$

$$(5.0) (50) = (2.0) V_2$$

$$(250.0) = (2.0) V_2$$

$$V_2 = 125.0 \text{ mL}$$

# Molarity By Dilution Worksheet Instructional Fair

**RJ Alexander**

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