

Workshop 1: Molarity, Molality and Mole Fraction

ANSWER KEY

In General Chemistry 2 there are some new ideas that depend on your understanding of 2 new units of concentration in addition to the ones we learned about in General Chemistry 1. These new units of concentration are *Molality (m)* and *Mole Fraction (x)*. This workshop is designed to help you review *Molarity (M)*, from General Chemistry 1) and then learn about what the new units of concentration (molality and mole fraction) mean.

Part 1. Let's begin by reviewing molarity:

Molarity (M) is the number of moles of a substance per 1 liter of solution. You learned this last semester in Topic 18. You also need to know how to use the dilution formula: $M_1V_1 = M_2V_2$. General Chemistry 2 is largely a course in aqueous solution chemistry and therefore *molarity* is probably the most important mole unit we will use all semester. You want to get *really good* at using molarity and of course you also need to remember naming and stoichiometry.

Before you write down your answers to these questions make sure that **EVERYONE** in your group is comfortable with the content that is being tested. Take a moment to discuss the questions in your group and when you are sure you all agree on the correct answers, write them down in the space provided below. You should ALL be able to do these types of questions in your sleep!!

Spend no more than 25 minutes on the questions in this section.

1. Consider a solution that is made by dissolving 6.5 g of calcium phosphate in 375 mL of water?

- a. Is calcium phosphate an ionic compound or a covalent compound? What is the chemical formula for calcium phosphate?

Ionic ☒

Covalent ☐

Both ☐

Chemical Formula: $\text{Ca}_3(\text{PO}_4)_2$

- b. When 1 molecule of calcium phosphate is added to water, what and how many ions are formed? Remember that ions are charged and so you **MUST** indicate the correct charge on each ion to receive credit for this question.

3 Ca^{2+} ions and 2 PO_4^{3-} ions

- c. How many moles of calcium phosphate are there in 6.5 g of calcium phosphate?

M_w of $\text{Ca}_3(\text{PO}_4)_2 = 310.18 \text{ g/mol}$

Moles of $\text{Ca}_3(\text{PO}_4)_2 = \text{mass}/M_w = 6.5 \text{ g}/(310.18 \text{ g/mol}) = 0.021 \text{ moles}$

- d. Finally, what is the molarity of the solution?

Molarity = moles/liter = $0.021 \text{ mol}/(0.375 \text{ L}) = 0.056 \text{ M}$

Molarity And Molality Notes And Practice Answers

H Kauffman



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Table of Contents Molarity And Molality Notes And Practice Answers

1. Understanding the eBook Molarity And Molality Notes And Practice Answers
 - The Rise of Digital Reading Molarity And Molality Notes And Practice Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity And Molality Notes And Practice Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity And Molality Notes And Practice Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity And Molality Notes And Practice Answers
 - Personalized Recommendations
 - Molarity And Molality Notes And Practice Answers User Reviews and Ratings

- Molarity And Molality Notes And Practice Answers and Bestseller Lists
- 5. Accessing Molarity And Molality Notes And Practice Answers Free and Paid eBooks
 - Molarity And Molality Notes And Practice Answers Public Domain eBooks
 - Molarity And Molality Notes And Practice Answers eBook Subscription Services
 - Molarity And Molality Notes And Practice Answers Budget-Friendly Options
- 6. Navigating Molarity And Molality Notes And Practice Answers eBook Formats
 - ePub, PDF, MOBI, and More
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 - Molarity And Molality Notes And Practice Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity And Molality Notes And Practice Answers
 - Highlighting and Note-Taking Molarity And Molality Notes And Practice Answers
 - Interactive Elements Molarity And Molality Notes And Practice Answers
- 8. Staying Engaged with Molarity And Molality Notes And Practice Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity And Molality Notes And Practice Answers
- 9. Balancing eBooks and Physical Books Molarity And Molality Notes And Practice Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity And Molality Notes And Practice Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Molarity And Molality Notes And Practice Answers
 - Setting Reading Goals Molarity And Molality Notes And Practice Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molarity And Molality Notes And Practice Answers
 - Fact-Checking eBook Content of Molarity And Molality Notes And Practice Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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