

Molar Mass Practice 1.0 – Answer Key**Objectives:** - Calculate molar mass using the periodic table**Directions:** Find the molar mass of the following atoms, molecules, or formula units. Round masses to the hundredths place. Don't forget your units! You must show your work!

1. Ti $\boxed{79.87 \text{ g/mol}}$



$$2 \times 12.01 + 4 \times 1.01 + 32.07 = \boxed{76.17 \text{ g/mol}}$$

2. Fe $\boxed{55.85 \text{ g/mol}}$



$$\text{C: } 2 \times 12.01 = 24.02$$

$$\text{H: } 6 \times 1.01 = 6.06$$

$$\text{O: } 1 \times 16.00 = 16.00$$

$$\boxed{46.11 \text{ g/mol}}$$

3. CO_2
 $\text{C: } 1 \times 12.01 = 12.01$
 $\text{O: } 2 \times 16.00 = 32.00$
$$\boxed{44.01 \text{ g/mol}}$$



$$3\text{Mg: } 3 \times 24.31 = 72.93$$

$$\text{N: } 6 \times 14.01 = 84.06$$

$$\text{O: } 12 \times 16.00 = 192.00$$

$$\boxed{349.00 \text{ g/mol}}$$

4. LiCl
 $\text{Li: } 1 \times 6.94 = 6.94$
 $\text{Cl: } 1 \times 35.45 = 35.45$
$$\boxed{42.39 \text{ g/mol}}$$



$$2 \times 1.01 + 32.07 + 4 \times 16.00 = \boxed{98.09 \text{ g/mol}}$$

5. O_2
 $\text{O: } 2 \times 16.00 = 32.00$
$$\boxed{32.00 \text{ g/mol}}$$

13. $\text{Mg}_3(\text{PO}_4)_2$
 $3\text{Mg: } 3 \times 24.31 = 72.93$
 $\text{P: } 2 \times 30.97 = 61.94$
 $\text{O: } 8 \times 16.00 = 128.00$
$$\boxed{262.87 \text{ g/mol}}$$

6. Iron(III) oxide Fe_2O_3
$$2 \times 55.85 + 3 \times 16.00 = \boxed{159.7 \text{ g/mol}}$$

14. Dihydrogen monoxide H_2O
$$2 \times 1.01 + 16.00 = \boxed{18.02 \text{ g/mol}}$$

7. $\text{HC}_2\text{H}_3\text{O}_2$
$$\boxed{75.07 \text{ g/mol}}$$

15. $\text{Zn}(\text{NO}_3)_2$
$$\boxed{297.4 \text{ g/mol}}$$

8. Lead(II) hydroxide $\text{Pb}(\text{OH})_2$
$$\boxed{241.2 \text{ g/mol}}$$

16. Aluminum sulfate $\text{Al}_2(\text{SO}_4)_3$
$$\boxed{342.15 \text{ g/mol}}$$

Molar Mass Practice Answer Key

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