

- (i) It depends on the concentration of a non-electrolyte solute in a solution and does not depend on the nature of the solute molecules.
- (ii) It depends on a number of particles of electrolyte solute in a solution and does not depend on the nature of the solute particles.
- (iii) It depends on the concentration of a non-electrolyte solute in solution as well as on the nature of the solute molecules.
- (iv) It depends on the concentration of an electrolyte or nonelectrolyte solute in solution as well as on the nature of solute molecules.

Solution:

Option (i) and (ii) are the answers.

30. Van't Hoff factor i is given by the expression _____.

- (i) $i = \text{Normal molar mass} / \text{Abnormal molar mass}$
- (ii) $i = \text{Abnormal molar mass} / \text{Normal molar mass}$
- (iii) $i = \text{Observed colligative property} / \text{calculated colligative property}$
- (iv) $i = \text{Calculated colligative property} / \text{Observed colligative property}$

Solution:

Option (i) and (iii) are the answers.

31. Isotonic solutions must have the same _____.

- (i) solute
- (ii) density
- (iii) elevation in boiling point
- (iv) depression in freezing point

Solution:

Option (iii) and (iv) are the answers.

32. Which of the following binary mixtures will have the same composition in the liquid and vapour phase?

- (i) Benzene - Toluene
- (ii) Water-Nitric acid
- (iii) Water-Ethanol
- (iv) n -Hexane - n -Heptane

Solution:

Option (ii) and (iii) are the answers.

33. In isotonic solutions _____.

- (i) solute and solvent both are same.
- (ii) osmotic pressure is the same.
- (iii) solute and solvent may or may not be same.
- (iv) a solute is always the same solvent may be different.

Solution:

Option (ii) and (iii) are the answers.

34. For a binary ideal liquid solution, the variation in total vapour pressure versus the composition of the solution is given by which of the curves?

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