

1. Calculate the molar mass of the compound (just answer in scientific notation)

$$C_2H_5NO_2 \rightarrow \frac{1 \text{ mol } C_2H_5NO_2}{1 \text{ mol } C_2H_5NO_2} = \frac{1 \text{ mol } C_2H_5NO_2}{1 \text{ mol } C_2H_5NO_2} = 75.07 \text{ g/mol}$$

2. Calculate the molar mass.

$$C_2H_5NO_2 \rightarrow \frac{1 \text{ mol } C_2H_5NO_2}{1 \text{ mol } C_2H_5NO_2} = 75.07 \text{ g/mol}$$

3. Calculate the molar mass.

$$C_2H_5NO_2 \rightarrow \frac{1 \text{ mol } C_2H_5NO_2}{1 \text{ mol } C_2H_5NO_2} = 75.07 \text{ g/mol}$$

4. Calculate using molar mass (just answer in scientific notation)

$$C_2H_5NO_2 \rightarrow \frac{1 \text{ mol } C_2H_5NO_2}{1 \text{ mol } C_2H_5NO_2} = 75.07 \text{ g/mol}$$

5. What is the empirical formula of the compound?



6. What is the empirical formula of the compound?



7. What is the empirical formula of the compound?



8. Name the compound (just answer in scientific notation)



9. Name the compound (just answer in scientific notation)



Modern Chemistry Mixed Review Chapteranswers

Patrick Vollmar



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