

Calculate the partial pressure of nitrogen gas, the moles of dry gas present, and the total pressure in the system. Show your work.

$$PV = nRT \rightarrow P_{\text{N}_2} = nRT/V \rightarrow (0.050\text{mol} \times 0.08206\text{L}\cdot\text{atm}/\text{K}\cdot\text{mol} \times 300\text{K})/2.0\text{L} =$$

$$P_{\text{N}_2} = 0.632\text{atm}$$

$$\rightarrow n_{\text{dry}} = PV/RT \rightarrow (0.9\text{atm} \times 2.0\text{L})/(0.08206\text{L}\cdot\text{atm}/\text{K}\cdot\text{mol} \times 300\text{K}) =$$

$$n_{\text{dry}} = 0.072\text{mol}$$

$$P_{\text{tot}} = P_{\text{N}_2} + P_{\text{H}_2\text{O}}$$

$$= 0.9\text{atm} + 0.632\text{atm} = P_{\text{tot}} = 1.54\text{atm}$$

7. Calculate the volume of hydrogen gas produced at 4.50 atm and -30°C by the reaction of 26.5g of calcium with excess hydrobromic acid according to the unbalanced equation. Show your work.



Student

$$PV = nRT \rightarrow 26.5\text{g Ca} \times (1\text{mol Ca}/40.08\text{g Ca}) \times (1\text{mol H}_2/1\text{mol Ca}) = 0.663\text{mol H}_2$$

$$V = nRT/P = (0.663\text{mol} \times 0.08206\text{L}\cdot\text{atm}/\text{K}\cdot\text{mol} \times 241\text{K})/4.50\text{atm} = V = 2.94\text{L}$$

8. Hydrogen Sulfide reacts with sulfur dioxide to produce water and sulfur. If 3.2g of sulfur was produced under the following conditions: 6.0L @ 750 Torr, what is the temperature of the hydrogen sulfide gas, in °C?



$$3.2\text{g S} \times (1\text{mol S}/32.07\text{g S}) \times (2\text{mol H}_2\text{S}/3\text{mol S}) = 0.067\text{mol H}_2\text{S}$$

$$PV = nRT \rightarrow T = PV/nR \rightarrow (0.937\text{atm} \times 6.0\text{L})/(0.067\text{mol} \times 0.08206$$

$$\text{L}\cdot\text{atm}/\text{K}\cdot\text{mol}) = 1077.12\text{K}$$

$$= 804^\circ\text{C}$$

Student - Student will apply Kinetic Molecular Theory when describing a novel, real-world situation in terms of collisions and motions of particles.

9. In the summer, the tire pressure sensor in your car alerts you that your tires need to be checked. At the gas station, you find that the tire pressure is too high inside the tire, however you have not added additional gas. Using the Kinetic Molecular Theory, explain why this occurs, making sure to refer to the motion of the gas molecules.

Since the summer is known for high temps, a high T would cause for the motion of gas molecules inside the tire to increase. With the increase in motion comes an increase in the number of collisions. With more collisions occurring, this results in an increase in pressure inside the tire.

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1. Understanding the eBook Modern Chemistry Gases Section 2 Review Answers
 - The Rise of Digital Reading Modern Chemistry Gases Section 2 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Gases Section 2 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Gases Section 2 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Gases Section 2 Review Answers
 - Personalized Recommendations
 - Modern Chemistry Gases Section 2 Review Answers User Reviews and Ratings
 - Modern Chemistry Gases Section 2 Review Answers and Bestseller Lists
5. Accessing Modern Chemistry Gases Section 2 Review Answers Free and Paid eBooks
 - Modern Chemistry Gases Section 2 Review Answers Public Domain eBooks
 - Modern Chemistry Gases Section 2 Review Answers eBook Subscription Services
 - Modern Chemistry Gases Section 2 Review Answers Budget-Friendly Options

6. Navigating Modern Chemistry Gases Section 2 Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Gases Section 2 Review Answers Compatibility with Devices
 - Modern Chemistry Gases Section 2 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Gases Section 2 Review Answers
 - Highlighting and Note-Taking Modern Chemistry Gases Section 2 Review Answers
 - Interactive Elements Modern Chemistry Gases Section 2 Review Answers
8. Staying Engaged with Modern Chemistry Gases Section 2 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Gases Section 2 Review Answers
9. Balancing eBooks and Physical Books Modern Chemistry Gases Section 2 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Gases Section 2 Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Gases Section 2 Review Answers
 - Setting Reading Goals Modern Chemistry Gases Section 2 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Gases Section 2 Review Answers
 - Fact-Checking eBook Content of Modern Chemistry Gases Section 2 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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- Interactive and Gamified eBooks

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