

Holt Modern Chemistry Review

CHAPTER 5: THE PERIODIC LAW

*The following pages contain the bulk (but not all) of the information for the chapter 5 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test. ***** Test items will be recall, examples, and/or application of this content. ******

OUTCOMES

- Collaborate with peer(s) to understand chemistry content (C C)
- Communicate chemistry content to teacher and peer(s) (E C)
- 5.1: Describe the organization of the modern Periodic table; distinguish between periods and groups (T & R)
- 5.2: Use the Periodic Table to obtain information about the properties (physical and chemical), ionization energy, and atomic sizes of the elements (F & PK)
- 5.2: Identify common non-metals, metalloids, and noble gases; identify the representative groups and name them (groups 1, 2, 13, 14, 15, 16, 17, 18) (T & R)
- 5.3: Compare periodic trends of atomic radii, ionization energy, and electronegativity (T & R)

5.1: HISTORY OF THE PERIODIC TABLE

- **Vocabulary**
 - **periodic law** -- the law that states that the repeating chemical and physical properties of elements change periodically with the atomic numbers of the elements
 - **periodic table** -- an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column
 - **lanthanides** -- a member of the rare-earth series of elements, whose atomic numbers range from 58 (cerium) to 71 (lutetium)
 - **actinides** -- any of the series of heavy radioactive elements that extends from thorium (atomic number 90) through lawrencium (atomic number 103) on the periodic table
- **Chapter Highlights**
 - The periodic law states that the physical and chemical properties of the elements are periodic functions of their atomic numbers.
 - The periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family.
 - The columns in the periodic table are referred to as groups.

5.2: ELECTRON CONFIGURATION AND THE PERIODIC TABLE

- **Vocabulary**
 - **alkali metal** -- one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium)
 - **alkaline-earth metal** -- one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
 - **transition element** -- one of the metals that can use the inner shell before using the outer shell to bond
 - **main-group element** -- an element in the s-block or p-block of the periodic table
 - **halogen** -- one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine); halogens combine with most metals to form salts
- **Chapter Highlights**
 - The rows in the periodic table are called periods.
 - Many chemical properties of the elements can be explained by the configurations of the elements' outermost (valence) electrons.
 - The noble gases exhibit unique chemical stability because their highest occupied levels have an octet (8) of electrons, $?s^2?p^6$ (with the exception of helium, whose stability arises from its highest occupied level being completely filled with two electrons, $1s^2$).

Modern Chemistry Chapter 5 2 Review The Answer Key

Lei Shi

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Table of Contents Modern Chemistry Chapter 5 2 Review The Answer Key

1. Understanding the eBook Modern Chemistry Chapter 5 2 Review The Answer Key
 - The Rise of Digital Reading Modern Chemistry Chapter 5 2 Review The Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 5 2 Review The Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Chapter 5 2 Review The Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter 5 2 Review The Answer Key
 - Personalized Recommendations
 - Modern Chemistry Chapter 5 2 Review The Answer Key User Reviews and Ratings
 - Modern Chemistry Chapter 5 2 Review The Answer Key and Bestseller Lists
5. Accessing Modern Chemistry Chapter 5 2 Review The Answer Key Free and Paid eBooks
 - Modern Chemistry Chapter 5 2 Review The Answer Key Public Domain eBooks
 - Modern Chemistry Chapter 5 2 Review The Answer Key eBook Subscription Services
 - Modern Chemistry Chapter 5 2 Review The Answer Key Budget-Friendly Options

6. Navigating Modern Chemistry Chapter 5 2 Review The Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Chapter 5 2 Review The Answer Key Compatibility with Devices
 - Modern Chemistry Chapter 5 2 Review The Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 5 2 Review The Answer Key
 - Highlighting and Note-Taking Modern Chemistry Chapter 5 2 Review The Answer Key
 - Interactive Elements Modern Chemistry Chapter 5 2 Review The Answer Key
8. Staying Engaged with Modern Chemistry Chapter 5 2 Review The Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Chapter 5 2 Review The Answer Key
9. Balancing eBooks and Physical Books Modern Chemistry Chapter 5 2 Review The Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Chapter 5 2 Review The Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Chapter 5 2 Review The Answer Key
 - Setting Reading Goals Modern Chemistry Chapter 5 2 Review The Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Chapter 5 2 Review The Answer Key
 - Fact-Checking eBook Content of Modern Chemistry Chapter 5 2 Review The Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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