#### Honors Chemistry Chapter 5 Review

#### Section 5.1: The Nature of Energy

- Kinetic and potential energy
- Electrostatic potential energy
  - o When is it positive and negative?
- · Units of energy
- · Systems and surroundings
  - o Open, closed, isolated systems
- Transferring energy
  - Work and Heat

#### Section 5.2: The first law of thermodynamics

- Internal energy
  - Know the direction of E, q, & w
- Endothermic & endothermic processes
- State functions

#### Section 5.3: Enthalpy

Pressure volume work

#### Section 5.4: Enthalpies of reaction

Calculate the energy absorbed or released in a reaction

#### Section 5.5: Calorimetry

- Heat capacity / molar heat capacity
- Constant pressure calorimetry
  - Calculations involving the heat gained or lost during a reaction
- For a given process at constant pressure, ΔH is negative. This means that the process is \_\_\_\_\_\_\_.
  - A) endothermic
  - B) equithermic
  - C) exothermic
  - D) a state function
  - E) energy
- 2. Which of the following statements is false?
  - A) Internal energy is a state function.
  - B) Enthalpy is an intensive property.
  - C) The enthalpy change for a reaction is equal in magnitude, but opposite in sign, to the enthalpy change for the reverse reaction.
  - The enthalpy change for a reaction depends on the state of the reactants and products.
  - E) The enthalpy of a reaction is equal to the heat of the reaction at constant pressure.

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