

Chapter 15

Oxidation-Reduction Reactions

SECTION 1

Transfer electrons around the following equations for the species indicated.

1. _____ Add the following species to complete each reaction correctly.
- For $\text{Cr}^{3+} \rightarrow \text{Cr}^{6+}$ _____
- For $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+}$ _____
- For $\text{Mg}^{2+} \rightarrow \text{Mg}^{0}$ _____

Answers:

2. Assign oxidation numbers to each element in the indicated species as they follow:

1. _____ a. H_2SO_4 for S
b. H_2SO_4 for O
c. H_2SO_4 for H
d. H_2SO_4 for S
e. H_2SO_4 for H
f. H_2SO_4 for O
g. H_2SO_4 for S
h. H_2SO_4 for O

3. In each of the following half-reactions, determine the electron:

- a. _____ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
b. _____ $\text{Fe}^{3+} \rightarrow \text{Fe}^{2+} + e^-$
c. _____ $\text{Fe}^{3+} \rightarrow \text{Fe}^{2+} + e^-$

Answers: _____ d. $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$ (oxidation)
e. $\text{Fe}^{3+} \rightarrow \text{Fe}^{2+} + e^-$ (reduction)

4. Name compounds other than those listed in Figure 15.1 of the text, but the following:

Answers:

1. _____ a. Fe^{2+} is a transition metal ion in the oxidation state.
b. Fe^{3+} is a transition metal ion.
c. Fe^{2+} is a transition metal ion in the oxidation state.
d. Fe^{3+} is a transition metal ion in the oxidation state.

Modern Chemistry Ch 19 Section Review 1

Richard Bailey



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Table of Contents Modern Chemistry Ch 19 Sectin Review 1

1. Understanding the eBook Modern Chemistry Ch 19 Sectin Review 1
 - The Rise of Digital Reading Modern Chemistry Ch 19 Sectin Review 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Ch 19 Sectin Review 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Ch 19 Sectin Review 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Ch 19 Sectin Review 1
 - Personalized Recommendations
 - Modern Chemistry Ch 19 Sectin Review 1 User Reviews and Ratings
 - Modern Chemistry Ch 19 Sectin Review 1 and Bestseller Lists
5. Accessing Modern Chemistry Ch 19 Sectin Review 1 Free and Paid eBooks
 - Modern Chemistry Ch 19 Sectin Review 1 Public Domain eBooks
 - Modern Chemistry Ch 19 Sectin Review 1 eBook Subscription Services
 - Modern Chemistry Ch 19 Sectin Review 1 Budget-Friendly Options

6. Navigating Modern Chemistry Ch 19 Sectin Review 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Ch 19 Sectin Review 1 Compatibility with Devices
 - Modern Chemistry Ch 19 Sectin Review 1 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Ch 19 Sectin Review 1
 - Highlighting and Note-Taking Modern Chemistry Ch 19 Sectin Review 1
 - Interactive Elements Modern Chemistry Ch 19 Sectin Review 1
8. Staying Engaged with Modern Chemistry Ch 19 Sectin Review 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Ch 19 Sectin Review 1
9. Balancing eBooks and Physical Books Modern Chemistry Ch 19 Sectin Review 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Ch 19 Sectin Review 1
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Ch 19 Sectin Review 1
 - Setting Reading Goals Modern Chemistry Ch 19 Sectin Review 1
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Ch 19 Sectin Review 1
 - Fact-Checking eBook Content of Modern Chemistry Ch 19 Sectin Review 1
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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