



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the reaction:

What is Equilibrium? Dynamic Equilibrium: A chemical reaction in which the concentrations of the reactants and products remain constant over time. For a reversible reaction, the equilibrium constant (K_c) is a measure of the extent to which the reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. Equilibrium Constant (K_c) For the reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant (K_c) For the reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Graphical Representation of Equilibrium Concentration Time Reactants $K_c < 1$ ($[A]_0 > [B]_0$) Products $K_c > 1$ ($[A]_0 < [B]_0$)								
Equilibrium Constant (K_c) For the reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant (K_c) <table><tr><th>Equilibrium Constant (K_c)</th><th>Meaning</th></tr><tr><td>$K_c > 1$</td><td>Concentration of products is greater than concentration of reactants.</td></tr><tr><td>$K_c < 1$</td><td>Concentration of reactants is greater than concentration of products.</td></tr><tr><td>$K_c = 1$</td><td>Concentration of reactants is equal to concentration of products.</td></tr></table>		Equilibrium Constant (K_c)	Meaning	$K_c > 1$	Concentration of products is greater than concentration of reactants.	$K_c < 1$	Concentration of reactants is greater than concentration of products.	$K_c = 1$	Concentration of reactants is equal to concentration of products.
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Le Chatelier's Principle: the effect of changing conditions on equilibrium										
Concentration	1. Increasing concentration of a reactant or product shifts the equilibrium position.	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing concentration of N_2 or H_2 shifts the equilibrium position to the right, producing more NH_3.								
Pressure	2. Increasing pressure shifts the equilibrium position towards the side with fewer moles of gas.	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing pressure shifts the equilibrium position to the right, producing more NH_3.								
Temperature	3. Increasing temperature shifts the equilibrium position towards the endothermic reaction.	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing temperature shifts the equilibrium position to the left, producing less NH_3.								

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