

Assessment #3 Review

Name: _____

You must show all work for credit.

Due Date: _____ (No late packets will be accepted)

I. Balance Equations*Balance the equations*

- $\text{Mg} + \underline{2} \text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$
- $\text{Mg} + \underline{2} \text{NaBr} \rightarrow \text{MgBr}_2 + \underline{2} \text{Na}$
- $\text{C}_3\text{H}_8 + \underline{5} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{4} \text{H}_2\text{O}$
- $\underline{4} \text{P} + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$
- $\text{Fe}_2(\text{SO}_4)_3 + \underline{6} \text{KOH} \rightarrow \underline{3} \text{K}_2\text{SO}_4 + \underline{2} \text{Fe}(\text{OH})_3$
- $\text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2 \rightarrow \text{Fe} + \underline{3} \text{H}_2\text{O}$

7. How many moles of hydrogen gas react with 1 mole of Fe_2O_3 in reaction #6? 3

Write the balanced equation for problems # 8 and 9:

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.
 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$

2 N on left, 2 N on right, 6 H on left, 6 H on right, 2 on left, 2 on right, so no net created or lost

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

2 N, 6 H, 2 on left, 2 on right

12. Using the balanced equation shown in problem #10, list the products of the reaction.

2NH₃

13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

N₂, 3H₂

II. Types of Reactions*Directions: Identify the type of reaction.*

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



Modern Chemistry Semester Exam Review Answers

Anjali Khirwadkar



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1. Understanding the eBook Modern Chemistry Semester Exam Review Answers
 - The Rise of Digital Reading Modern Chemistry Semester Exam Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Semester Exam Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Semester Exam Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Semester Exam Review Answers
 - Personalized Recommendations
 - Modern Chemistry Semester Exam Review Answers User Reviews and Ratings

- Modern Chemistry Semester Exam Review Answers and Bestseller Lists
- 5. Accessing Modern Chemistry Semester Exam Review Answers Free and Paid eBooks
 - Modern Chemistry Semester Exam Review Answers Public Domain eBooks
 - Modern Chemistry Semester Exam Review Answers eBook Subscription Services
 - Modern Chemistry Semester Exam Review Answers Budget-Friendly Options
- 6. Navigating Modern Chemistry Semester Exam Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Semester Exam Review Answers Compatibility with Devices
 - Modern Chemistry Semester Exam Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Semester Exam Review Answers
 - Highlighting and Note-Taking Modern Chemistry Semester Exam Review Answers
 - Interactive Elements Modern Chemistry Semester Exam Review Answers
- 8. Staying Engaged with Modern Chemistry Semester Exam Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Semester Exam Review Answers
- 9. Balancing eBooks and Physical Books Modern Chemistry Semester Exam Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Semester Exam Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Semester Exam Review Answers
 - Setting Reading Goals Modern Chemistry Semester Exam Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Semester Exam Review Answers
 - Fact-Checking eBook Content of Modern Chemistry Semester Exam Review Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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