

CH235-2E
ORGANIC CHEMISTRY I
Quiz 9

Concept and Rate Equation

1. Which of the following reactions corresponds to a *substitution*?
 - a. *tert*-butanol $\rightarrow$ *tert*-butyl chloride
 - b. *tert*-butanol $\rightarrow$ 2-methylpropene
 - c. 3,3-dimethyl-2-butanol $\rightarrow$ 2,3-dimethyl-2-butene
 - d. cyclohexene $\rightarrow$ 1,2-dichlorocyclohexane
2. Which of the following statements related to S_N1 reactions is *not* true?
 - a. The heterolysis of a bond between atoms which do not bear formal charges always produces a cation and an anion
 - b. The charged carbon atom of a carbocation has a complete octet of valence shell electrons
 - c. Carbocations are Lewis acids
 - d. Nucleophiles seek centers of low electron density
3. What is the equation for the rate of formation of *tert*-butyl alcohol from the reaction of *tert*-butyl bromide (*t*-BuBr) with water by an S_N1 mechanism?
 - a. $\text{Rate} = k [\textit{t}\text{-BuBr}]$
 - b. $\text{Rate} = k [\textit{t}\text{-BuBr}][\text{H}_2\text{O}]$
 - c. $\text{Rate} = k [\text{H}_2\text{O}]$
 - d. $\text{Rate} = k [\textit{t}\text{-BuBr}]^2$
4. What is the equation for the rate of formation of 2-methoxypropane, $(\text{CH}_3)_2\text{CH}(\text{OCH}_3)\text{CH}_3$, from the reaction of 2-bromopropane (*i*-PrBr) with sodium methoxide (NaOCH_3)?
 - a. $\text{Rate} = k [\textit{i}\text{-PrBr}]$
 - b. $\text{Rate} = k [\textit{i}\text{-PrBr}]^2$
 - c. $\text{Rate} = k [\text{NaOCH}_3]$
 - d. $\text{Rate} = k [\textit{i}\text{-PrBr}][\text{NaOCH}_3]$
5. The reaction of methyl iodide with sodium azide, NaN_3 , proceeds by an S_N2 mechanism. What is the effect of doubling the concentration of NaN_3 on the rate of the reaction?
 - a. the rate remains the same
 - b. the rate decreases by a factor of 2
 - c. the rate increases by a factor of 2
 - d. the rate increases by a factor of 4
6. The reaction of *tert*-butyl bromide, $(\text{CH}_3)_3\text{CBr}$, with methanol in an inert solvent proceeds by an S_N1 mechanism to give *tert*-butyl methyl ether, $(\text{CH}_3)_3\text{COCH}_3$. What is the effect of doubling the concentration of methanol on the rate of the reaction?
 - a. the rate remains the same
 - b. the rate decreases by a factor of 2
 - c. the rate increases by a factor of 2
 - d. the rate increases by a factor of 4

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Oswaal Editorial Board



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