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CHAPTER 2 – The Chemistry of Life

CHAPTER OUTLINE

2.1 Atoms Make Up All Matter

- A. Elements Are Fundamental Types of Matter
- B. Atoms Are Particles of Elements
- C. Isotopes Have Different Numbers of Neutrons

2.2 Chemical Bonds Link Atoms

- A. Electrons Determine Bonding
- B. In an Ionic Bond, One Atom Transfers Electrons to Another Atom
- C. In a Covalent Bond, Atoms Share Electrons
- D. Partial Charges on Polar Molecules Create Hydrogen Bonds

2.3 Water Is Essential to Life

- A. Water Is Cohesive and Adhesive
- B. Many Substances Dissolve in Water
- C. Water Regulates Temperature
- D. Water Expands as It Freezes
- E. Water Participates in Life's Chemical Reactions

2.4 Cells Have an Optimum pH

2.5 Cells Contain Four Major Types of Organic Molecules

- A. Large Organic Molecules Are Composed of Smaller Subunits
- B. Carbohydrates Include Simple Sugars and Polysaccharides
- C. Proteins Are Complex and Highly Versatile
- D. Nucleic Acids Store and Transmit Genetic Information
- E. Lipids Are Hydrophobic and Energy-Rich

LEARNING OUTCOMES

- 02.00.01 Explain the relationship between chemistry and biology.
- 02.01.01 Identify the most important elements in living organisms.
- 02.01.02 Describe the structure of atoms.
- 02.02.01 Compare and contrast the different types of bonds.
- 02.02.02 Differentiate between atoms and molecules.

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