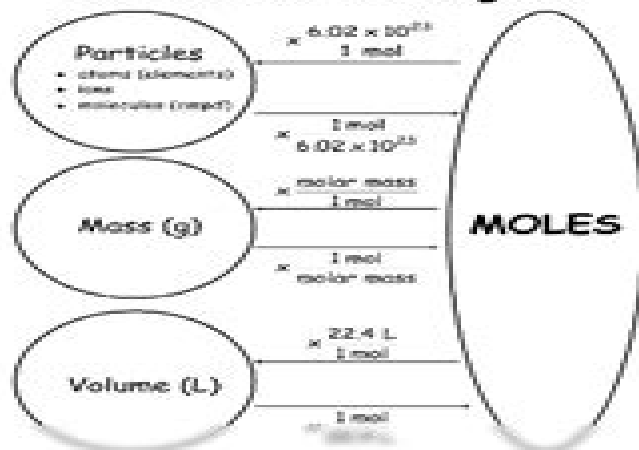


# Honors Chemistry B- Final Exam Formula Sheet

Note: You will not use all of these formulas for this test

Temperature Formulas:  
 $^{\circ}\text{C} = (^{\circ}\text{F} - 32) \cdot \frac{5}{9}$      $^{\circ}\text{F} = \frac{9}{5}^{\circ}\text{C} + 32$      $\text{K} = ^{\circ}\text{C} + 273$

## Mole Conversion Diagram



$1 \text{ atm} = 760 \text{ mm Hg} = 101.3 \text{ kPa}$

$P_1 + P_2 = P_3 + P_4$

$\frac{Q}{m} = \frac{Q}{m}$

$\frac{Q}{m} = \frac{Q}{m}$

$\Delta \frac{Q}{m} = \Delta \frac{Q}{m} - \Delta \frac{Q}{m}$

$P_1 + V_1 + A_1 + L_1 + T_1 = P_2 + V_2 + A_2 + L_2 + T_2$

$P_{\text{total}} = P_1 + P_2 + P_3 + \dots$

$\frac{Q}{m} = \frac{Q}{m}$

$\text{Pressure} = \frac{\text{Force}}{\text{Area}} = \frac{m \cdot g}{A} = \frac{m \cdot g}{\pi r^2}$

$M_1 \times V_1 = M_2 \times V_2$

Percent by Mass =  $\frac{\text{Mass of Solute}}{\text{Mass of Solution}} \times 100$

Percent by Volume =  $\frac{\text{Volume of Solute}}{\text{Volume of Solution}} \times 100$

Mole fractions:  $X_A = \frac{n_A}{n_A + n_B}$      $X_B = \frac{n_B}{n_A + n_B}$

1 Calorie = 1 kilocalorie = 1000 calories

$Q = mc\Delta T$

$q_{\text{sys}} = \Delta H = -q_{\text{surr}} = -m \times C \times \Delta T$

$\Delta H^{\circ} = \Delta H_f^{\circ}(\text{products}) - \Delta H_f^{\circ}(\text{reactants})$

$K_c = \frac{[\text{C}]^c \cdot [\text{D}]^d}{[\text{A}]^a \cdot [\text{B}]^b}$

$\Delta G = \Delta H - T\Delta S$

$\ln K = \frac{-\Delta G^{\circ}}{RT}$

$K_w = [\text{H}^+] \cdot [\text{OH}^-] = 1.0 \times 10^{-14}$

$\text{pH} = -\log[\text{H}^+]$

$\text{pOH} = -\log[\text{OH}^-]$

$K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]}$

# High School Chemistry Honors Study Guide Final

**R Barnett**



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








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