

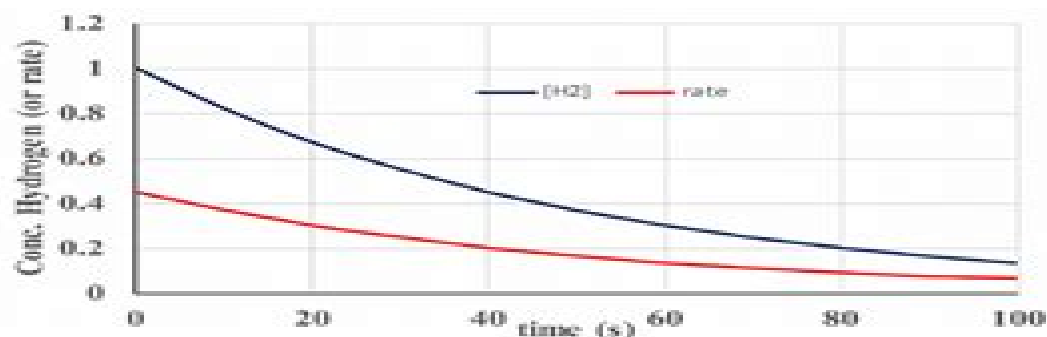
Chapter 15

Chemical Kinetics

1. The rate of chemical reaction – the change in the amounts of reagents (or products) divided by the change in time.

For the reaction of hydrogen and iodine, $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightarrow 2 \text{HI}(\text{g})$, the reaction rate is defined as

Rate = $-\frac{\Delta[\text{H}_2]}{\Delta t} = -\frac{\Delta[\text{I}_2]}{\Delta t} = +\frac{1}{2} \frac{\Delta[\text{HI}]}{\Delta t}$ Note that reaction rate is a positive number, therefore, negative sign must be added when reagents are used (because their concentrations decrease) and positive sign when products are used (because concentrations of products increase). Factor $\frac{1}{2}$ is due to stoichiometry because two moles of HI are made from one mole of H_2 and I_2 . Concentration of hydrogen and the reaction rate (magnified 25 times) for the above reaction, is shown below



2. The instantaneous reaction rate - this is the slope of the concentration curve plotted against the reaction time. For general reaction: $a\text{A} + b\text{B} \rightarrow c\text{C} + d\text{D}$, the reaction rate is defined as:
Rate = $-\frac{1}{a} \frac{\Delta[\text{A}]}{\Delta t} = -\frac{1}{b} \frac{\Delta[\text{B}]}{\Delta t} = +\frac{1}{c} \frac{\Delta[\text{C}]}{\Delta t} = +\frac{1}{d} \frac{\Delta[\text{D}]}{\Delta t}$
3. Measuring reaction rates $\rightarrow$ determine concentration of reagents or products as a function of time, using suitable analytical techniques, for example, absorption of light or chromatography.
4. The rate laws: The reaction rate depends on concentration of reagents. Let's consider reaction $\text{A} \rightarrow \text{products}$ (for example, decomposition of sucrose into glucose and

General Chemistry Ii Chapter Chemical Kinetics Answers

Annelies Wilder-Smith



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