

**Review for the Full Year General Chemistry  
American Chemical Society Exam**

**(a) Thermodynamics, (b) Transition Metals, (c) Redox and Electrochemistry, (d) Nuclear and (e) Organic**

**THERMODYNAMICS**

84. A 10.0 g sample of silver is heated to 100.0 °C and then added to 20.0 g of water at 23.0 °C in an insulated calorimeter. At thermal equilibrium the temperature of the system was measured as 25.0 °C. What is the specific heat of silver? [specific heat =  $4.2 \text{ J g}^{-1} \text{ }^{\circ}\text{C}^{-1}$ ]

- (A) 0.11 J/g K (C) 17 J/g K  
(B) 0.22 J/g K (D) 34 J/g K

85. What is  $\Delta H$  in kJ for the reaction:



- (A) -1830 kJ (B) -2284 kJ  
(C) -2510 kJ (D) -1605 kJ

86. Given these equations:



calculate the heat of formation of  $\text{SO}_3(\text{g})$ .

- (A) -500 kJ mol<sup>-1</sup> (C) +100 kJ mol<sup>-1</sup>  
(B) -400 kJ mol<sup>-1</sup> (D) +200 kJ mol<sup>-1</sup>

87. Given:



How much energy is produced in the combustion of 28 g of CO? [Atomic Masses: C 12.0 g mol<sup>-1</sup>; O 16.0 g mol<sup>-1</sup>]

- (A) 88 kJ (B) 109 kJ  
(C) 175 kJ (D) 284 kJ

88. Calculate the value of  $\Delta H$  (in kJ mol<sup>-1</sup>) for the reaction:



| Bond Energies (kJ mol <sup>-1</sup> ) |                          |
|---------------------------------------|--------------------------|
| H-H                                   | 435                      |
| N=N                                   | 945 (in N <sub>2</sub> ) |
| N-H                                   | 389                      |

- (A) 2340 kJ of heat absorbed  
(B) 213 kJ of heat absorbed  
(C) 2340 kJ of heat evolved  
(D) 83 kJ of heat evolved

89. Which change is likely to be accompanied by the greatest increase in entropy?



90. For which process is the entropy change per mole the largest at constant temperature?



91. In which process is entropy decreased?

- (A) dissolving sugar in water  
(B) expanding a gas  
(C) evaporating a liquid  
(D) freezing water

92. When  $\text{Al}_2\text{O}_3(\text{s})$  is formed from the elements at standard conditions, the values of  $\Delta H^{\circ}$  and  $\Delta G^{\circ}$  at 298 K are -1676 kJ mol<sup>-1</sup> and -1577 kJ mol<sup>-1</sup>, respectively. The standard entropy of formation per mole, in joules per degree, will be

- (A) -332 (B) -157 (C) -93.3  
(D) -0.0933 (E) +15.7

93. Vaporization of a liquid is an example of a process for which

- (A)  $\Delta H$ ,  $\Delta S$ , and  $\Delta G$  are positive at all temperatures.  
(B)  $\Delta H$  and  $\Delta S$  are positive.  
(C)  $\Delta G$  is negative at low temperatures, positive at high temperatures.  
(D)  $\Delta H = \Delta S$

94. A particular chemical reaction has a negative  $\Delta H$  and negative  $\Delta S$ . Which statement is correct?

- (A) The reaction is spontaneous at all temperatures.  
(B) The reaction is nonspontaneous at all temperatures.  
(C) The reaction becomes spontaneous as temperature increases.  
(D) The reaction becomes spontaneous as temperature decreases.

# General Chemistry Exam 4

**Didier Musso**



## **General Chemistry Exam 4:**

## General Chemistry Exam 4 Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the power of words has become more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such may be the essence of the book **General Chemistry Exam 4**, a literary masterpiece that delves deep into the significance of words and their impact on our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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### Table of Contents General Chemistry Exam 4

1. Understanding the eBook General Chemistry Exam 4
  - The Rise of Digital Reading General Chemistry Exam 4
  - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Exam 4
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook General Chemistry Exam 4
  - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Exam 4
  - Personalized Recommendations
  - General Chemistry Exam 4 User Reviews and Ratings
  - General Chemistry Exam 4 and Bestseller Lists
5. Accessing General Chemistry Exam 4 Free and Paid eBooks

- General Chemistry Exam 4 Public Domain eBooks
- General Chemistry Exam 4 eBook Subscription Services
- General Chemistry Exam 4 Budget-Friendly Options
- 6. Navigating General Chemistry Exam 4 eBook Formats
  - ePub, PDF, MOBI, and More
  - General Chemistry Exam 4 Compatibility with Devices
  - General Chemistry Exam 4 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of General Chemistry Exam 4
  - Highlighting and Note-Taking General Chemistry Exam 4
  - Interactive Elements General Chemistry Exam 4
- 8. Staying Engaged with General Chemistry Exam 4
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers General Chemistry Exam 4
- 9. Balancing eBooks and Physical Books General Chemistry Exam 4
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection General Chemistry Exam 4
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Exam 4
  - Setting Reading Goals General Chemistry Exam 4
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Exam 4
  - Fact-Checking eBook Content of General Chemistry Exam 4
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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