

Name Key

1. Hydrogen, H_2 , is used as rocket fuel. The hydrogen is burned in oxygen, O_2 , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H_2 , is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

$$\frac{99.2 \text{ mol } H_2}{2 \text{ mol } H_2} \times \frac{-484 \text{ kJ}}{1} = -24,006 \text{ kJ}$$

$-24,006 \text{ kJ}$
 ~~$-24,006 \text{ kJ}$~~

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is ΔH for the following reaction?

reverse so
change sign of
 ΔH out in 2
quantity



Exothermic

3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

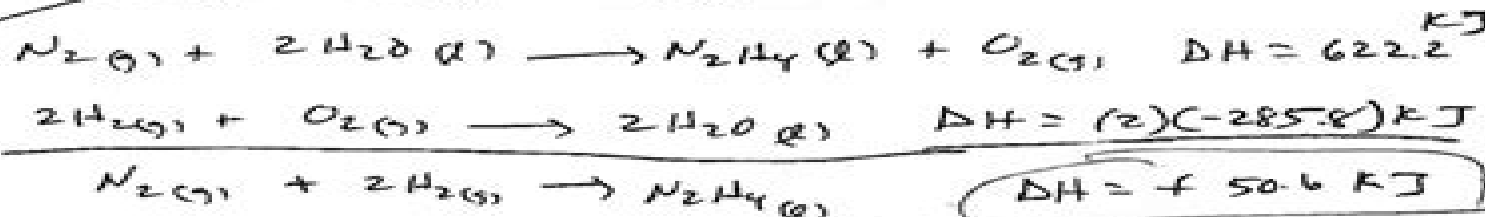
$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_8O} \times 4.5 \text{ mol } C_3H_8O = -1920.6 \text{ kcal}$$

4. The specific heat of iron is $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$. How much heat must be added to 5.2 g of iron at 25°C to raise its temperature to 225°C ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left(\frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N_2H_4 , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.



Reverse
double
quantity

General Chemistry Practice Problems With Answers

Mary Burnham, Carol Hurd



General Chemistry Practice Problems With Answers:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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The United States Catalog Mary Burnham, Carol Hurd, 1928

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Table of Contents General Chemistry Practice Problems With Answers

1. Understanding the eBook General Chemistry Practice Problems With Answers
 - The Rise of Digital Reading General Chemistry Practice Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Practice Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Practice Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Practice Problems With Answers
 - Personalized Recommendations
 - General Chemistry Practice Problems With Answers User Reviews and Ratings
 - General Chemistry Practice Problems With Answers and Bestseller Lists
5. Accessing General Chemistry Practice Problems With Answers Free and Paid eBooks
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 - General Chemistry Practice Problems With Answers eBook Subscription Services

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 - General Chemistry Practice Problems With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Practice Problems With Answers
 - Highlighting and Note-Taking General Chemistry Practice Problems With Answers
 - Interactive Elements General Chemistry Practice Problems With Answers
- 8. Staying Engaged with General Chemistry Practice Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Practice Problems With Answers
- 9. Balancing eBooks and Physical Books General Chemistry Practice Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Practice Problems With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Practice Problems With Answers
 - Setting Reading Goals General Chemistry Practice Problems With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Practice Problems With Answers
 - Fact-Checking eBook Content of General Chemistry Practice Problems With Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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