

## PRE-LAB QUESTIONS

1. Write names of the following monatomic ions.

Formula	Name
$K^+$	potassium ion
$Ni^{2+}$	nickel (II) ion
$Zn^{2+}$	zinc (II) ion
$Cr^{3+}$	chromium (III)
$F^-$	fluoride
$S^{2-}$	sulfide
$P^{3-}$	phosphide

2. Write the formulas of compounds formed by the following ions. For example, formula of compound with two
- $Na^+$
- ion and one
- $O^{2-}$
- ion is written as
- $Na_2O$
- . No charge should be written in the formulas of compounds.

a) compound with two  $Al^{3+}$  ions and three  $S^{2-}$  ions  $Al_2S_3$ b) compound with one  $Ni^{2+}$  ions and two  $Cl^-$  ions  $NiCl_2$ 

3. a) How many chloride ions (
- $Cl^-$
- ) are needed to form a compound with a barium ion (
- $Ba^{2+}$
- )?
- 2

(Hint: A compound should have a net charge of zero. Therefore, sum of charges of ions in a compound should be zero.)

b) What is the formula of compound formed containing  $Ba^{2+}$  ion and  $Cl^-$  ion?  
 $Cl_2Ba$

# Gcc Chm130ll Nuclear Chemistry Page 2

**Sabine Zange**





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