

CHM130LL
Lab 13 – Calorimetry

Name: _____ MEID: _____

Complete the following items by typing into the text box provided. The boxes will increase in size if additional space is necessary. Please be sure to save all your work as an MS Word document to submit properly.

Part 1: Determining the Specific Heat Capacity of Aluminum

A. Data Table 1

Complete the Data Table. (7 pts)

ALUMINUM METAL (Trial 1)	
Pre-weighed aluminum metal mass (g)	20.01 g
Initial temperature of cool water in the calorimeter (°C)	20.0°C
Initial temperature of aluminum metal (same as the boiling water in the p-10) (°C)	100.0°C
Final temperature of water/aluminum metal in calorimeter (°C)	24.0°C
ALUMINUM METAL (Trial 2)	
Pre-weighed aluminum metal mass (g) [Copy this from above table you are using the same metal sample for both trials.]	20.03 g
Initial temperature of cool water in the calorimeter (°C)	18.0°C
Initial temperature of aluminum metal (same as the boiling water in the p-10) (°C)	100.0°C
Final temperature of water/aluminum metal in calorimeter (°C)	21.8°C

B. Insert the following photo with your name and MEID showing in the photo.

Photo A: Insert a photo of your calorimeter showing your aluminum metal submerged in the water. The photo should be taken from above. (6 pts)

C. Calculations (Show all calculations with units on each number and the final answer rounded to the correct significant figures for full credit.)

- Calculate the mass of the water by using the water's density (assume it is 1.00 g/mL) and the volume of the water. Remember, density = mass/volume. (4 pts)

$$D = \frac{M}{V} \rightarrow M = D \times V \rightarrow M = \left[1.00 \frac{\text{g}}{\text{mL}} \right] (75 \text{ mL}) = 75 \text{ g water (Trial 1)}$$

$$D = \frac{M}{V} \rightarrow M = D \times V \rightarrow M = \left[1.00 \frac{\text{g}}{\text{mL}} \right] (75 \text{ mL}) = 75 \text{ g water (Trial 2)}$$

Gcc Chm130ll Nuclear Chemistry

**Walter D. Loveland, David J.
Morrissey, Glenn T. Seaborg**



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