

# **ACTIVITY: Molecular Polarity Computer Simulation (PhET)**

Go to: <https://phet.illinois.edu/> Click on the arrow, DOWNLOAD, OPEN and RUN the MOLECULE POLARITY program. Or google "phet molecule polarity" to find the simulation.

## **Introduction:**

In this atomic level simulation, you will investigate how atoms' **electronegativity** values affect the bonds they produce. When two atoms form a covalent bond, a pair of electrons is shared between atoms. **Electronegativity** is a measure of a single atom's ability to attract electrons shared in that bond. In this lab you will work at your own pace, to answer a number of questions. To begin, from what you've already learned about the protons and electrons in an atom, what would cause an atom to have a **high** electronegativity value?

Why might an atom have a low electronegativity value?

## **Procedure:**

**Two Atoms**

1. Turn on (check) all view options. **distill** out just the molecule with your mouse.
2. Take your time and investigate how the binary compound's bond behaves when the atom's electronegativity and orientation are changed. **Do not rush through this step.** Describe the bond formed between two atoms with **similar, low** electronegativities.

Describe the bond formed between two atoms with **similar, high** electronegativities.

Describe the bond formed between two atoms with **very different** electronegativities.

Describe (in your own words) what is meant by **partial charges**,  $\delta^-$  and  $\delta^+$ .

$\delta^-$  represents: \_\_\_\_\_

$\delta^+$  represents: \_\_\_\_\_

What happens when the electric field is applied to a very polar molecule?

Why do you think this is? \_\_\_\_\_

**View**

- ☒ Bond Dipoles **ON**
- ☒ Partial Charges
- ☒ Bond Orientation

**Surfaces**

- ☐ None
- ☐ Electrostatic Potential
- ☒ Electron Density

**Electric Field**

- ☐ On ☐ Off



# Molecule Polarity Phet Lab Key

**SA Adler**



**Molecule Polarity Phet Lab Key:**

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## **Table of Contents Molecule Polarity Phet Lab Key**

1. Understanding the eBook Molecule Polarity Phet Lab Key
  - The Rise of Digital Reading Molecule Polarity Phet Lab Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Molecule Polarity Phet Lab Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molecule Polarity Phet Lab Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molecule Polarity Phet Lab Key
  - Personalized Recommendations
  - Molecule Polarity Phet Lab Key User Reviews and Ratings
  - Molecule Polarity Phet Lab Key and Bestseller Lists
5. Accessing Molecule Polarity Phet Lab Key Free and Paid eBooks
  - Molecule Polarity Phet Lab Key Public Domain eBooks
  - Molecule Polarity Phet Lab Key eBook Subscription Services
  - Molecule Polarity Phet Lab Key Budget-Friendly Options
6. Navigating Molecule Polarity Phet Lab Key eBook Formats

- ePub, PDF, MOBI, and More
  - Molecule Polarity Phet Lab Key Compatibility with Devices
  - Molecule Polarity Phet Lab Key Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Molecule Polarity Phet Lab Key
    - Highlighting and Note-Taking Molecule Polarity Phet Lab Key
    - Interactive Elements Molecule Polarity Phet Lab Key
  8. Staying Engaged with Molecule Polarity Phet Lab Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Molecule Polarity Phet Lab Key
  9. Balancing eBooks and Physical Books Molecule Polarity Phet Lab Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Molecule Polarity Phet Lab Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Molecule Polarity Phet Lab Key
    - Setting Reading Goals Molecule Polarity Phet Lab Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Molecule Polarity Phet Lab Key
    - Fact-Checking eBook Content of Molecule Polarity Phet Lab Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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