

Unit 7 Test Review-Answer Key

1. What is the percent composition of Si_2O_6 ?

$$\begin{aligned} \bullet \text{ Si} &= (2 \times 28.09) = 56.18 \div 152.18 \times 100 = 36.9 \% \\ \bullet \text{ O} &= (6 \times 16) = 96 \div 152.18 \times 100 = 63.1 \% \\ \text{Total MM} &= 152.18 \end{aligned}$$

2. What is the percent composition of $\text{Ti}(\text{SO}_4)_2$?

$$\begin{aligned} \text{Ti} &= (1 \times 47.88) = 47.88 \div 240.02 \times 100 = 19.9 \% \\ \text{S} &= (2 \times 32.07) = 64.14 \div 240.02 \times 100 = 26.7 \% \\ \text{O} &= (8 \times 16) = 128 \div 240.02 \times 100 = 53.3 \% \\ \text{Total MM} &= 240.02 \end{aligned}$$

3. What is the percent composition of Hydrogen in $\text{C}_6\text{H}_{12}\text{O}_6$?

$$\begin{aligned} \text{C} &= (6 \times 12.01) = 72.06 \div 180.18 \times 100 \\ \text{H} &= (12 \times 1.01) = 12.12 \div 180.18 \times 100 = 6.7 \% \\ \text{O} &= (6 \times 16) = 96 \div 180.18 \times 100 \\ \text{Total MM} &= 180.18 \end{aligned}$$

4. An unknown compound is 43.4% sodium, 11.3% carbon, and 45.3% oxygen. What is the empirical formula?

$$\text{Na} = 43.4 \text{ g} \times \frac{1 \text{ mole}}{22.99 \text{ g}} = 1.89 \text{ moles} \div 0.94 = 2$$

$$\text{C} = 11.3 \text{ g} \times \frac{1 \text{ mole}}{12.01 \text{ g}} = 0.94 \text{ moles} \div 0.94 = 1 \quad \text{Empirical Formula} = \text{Na}_2\text{CO}_3$$

$$\text{O} = 45.3 \text{ g} \times \frac{1 \text{ mole}}{16 \text{ g}} = 2.93 \text{ moles} \div 0.94 = 3$$

5. Monosodium glutamate (MSG) is sometimes added to food to enhance flavor. Analysis determined this compound to be 35.5% C, 4.77% H, 8.28% N, 13.6% Na, and 37.9% O. What is the empirical formula for MSG?

$$\text{C} = 35.5 \text{ g} \times \frac{1 \text{ mole}}{12.01 \text{ g}} = 2.96 \text{ moles} \div 0.59 = 5.01 = 5$$

$$\text{H} = 4.77 \text{ g} \times \frac{1 \text{ mole}}{1.01 \text{ g}} = 4.7 \text{ moles} \div 0.59 = 7.96 = 8$$

$$\text{N} = 8.28 \text{ g} \times \frac{1 \text{ mole}}{14.01 \text{ g}} = 0.59 \text{ moles} \div 0.59 = 1$$

$$\text{Na} = 13.6 \text{ g} \times \frac{1 \text{ mole}}{22.99 \text{ g}} = 0.59 \text{ moles} \div 0.59 = 1$$

$$\text{O} = 37.9 \text{ g} \times \frac{1 \text{ mole}}{16 \text{ g}} = 2.37 \text{ moles} \div 0.59 = 4.01 = 4$$

$$\text{Empirical Formula} = \text{C}_5\text{H}_8\text{NNaO}_4$$

Mole Test Answer

Terry C. Jones



Mole Test Answer:

The Top Books of the Year Mole Test Answer The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous engrossing novels enthralling the hearts of readers worldwide. Lets delve into the realm of popular books, exploring the captivating narratives that have enthralled audiences this year. Mole Test Answer : Colleen Hoover's "It Ends with Us" This poignant tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover masterfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can prevail. Mole Test Answer : Taylor Jenkins Reid's "The Seven Husbands of Evelyn Hugo" This captivating historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reid's compelling storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Discover the Magic : Delia Owens' "Where the Crawdads Sing" This evocative coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These popular novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of engaging stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a masterful and gripping novel that will keep you wondering until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

<https://staging.conocer.cide.edu/book/book-search/Documents/Fish%20Pocket%20File.pdf>

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