

Module Overview

What are the coefficients in a chemical equation for combustion?

What?

A chemical equation represents what we see when we observe a reaction and predict products from reactants. Coefficients in formulas before them are necessary to conserve mass during a chemical reaction. Typically, when we balance the chemical equation we start by counting individual products. However, in real life the reaction represented by an equation occurs as unsegmented, constant volume. Most of what you see (large molecules) are the biggest building blocks shown in the equation. Some are the smallest chemical equations where they occur naturally. Equations often represent an end state for the reaction you will observe for different inputs chemical reaction rate for comparison.

Module 1 – A Chemical Reaction

$$10\text{H}_2\text{O}_2 + 10\text{H}_2\text{O}_2 \rightarrow 10\text{H}_2\text{O} + 10\text{O}_2$$

1. Consider the reaction in Module 1.

a. What are the coefficients for each of the following substances in the reaction?



b. Does your coefficient value in balance the equation in Module 1?

2. Consider each reaction below as a chemical reaction in Module 1.

a. What are the masses of reactants consumed and products made?

b. Based on the ratio of H_2O to O_2 in H_2O_2 , balance the reaction the lowest whole coefficient possible.

	H_2O_2 Reactants	H_2O Reactants	O_2 Products	Ratio H_2O to O_2 (Products)
How a single product from many reactants? What coefficient would be consumed or produced?				
If the reaction consumed more than double water, how many molecules would be consumed or produced?				
If the reaction consumed 100 times more water, molecules would be consumed or produced?				

Mole Ratios Worksheet Pogil

Lei Huang

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