Mole/Mole Ratio Problems

1. __N₂ + __H₂ → __NH₃

- a. How many moles of hydrogen are needed to completely react with two moles of nitrogen?
- b. How many moles of nitrogen trihydride can be produced with 5 moles of nitrogen?
- c. How many moles of nitrogen are needed to produce .5 moles of nitrogen trihydride?
- d. How many liters of nitrogen trihydride can be produced from 24 liters of nitrogen gas?
 - KCIO₃ → KCI + O₂
 - a. How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate?
 - b. How many moles of potassium chloride are produced by the decomposition of .75 moles of potassium chlorate?
 - c. How many moles of potassium chlorate are needed to produce 100 moles of oxygen?
 - d. How many liters of oxygen will the decomposition of 100 liters of potassium chlorate produce?

Mole Ratio Worksheet 17

O. Jerry Parker, Gary L. Breneman

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