

Molarity Problems

KEY

NOT IN ORDER

1. Sea water contains roughly 28.0 g of NaCl per liter. What is the molarity of sodium chloride in sea water?

$$\frac{28.0g \text{ NaCl}}{1} \times \frac{1 \text{ mol NaCl}}{58.44g} = \boxed{0.48 \text{ M}}$$

2. What is the molarity of 245.0 g of H_2SO_4 dissolved in 1.00 L of solution?

$$\frac{245.0g \text{ H}_2\text{SO}_4}{1} \times \frac{1 \text{ mol H}_2\text{SO}_4}{98.08g} = \boxed{2.5 \text{ M}}$$

3. What is the molarity of 5.30 g of Na_2CO_3 dissolved in 400.0 mL solution?

$$\frac{5.30g \text{ Na}_2\text{CO}_3}{1} \times \frac{1 \text{ mol Na}_2\text{CO}_3}{106g} = \frac{0.05 \text{ mol}}{.400 \text{ L}} = \boxed{0.125 \text{ M}}$$

4. What is the molarity of 5.00 g of NaOH in 750.0 mL of solution?

$$\frac{5.00g \text{ NaOH}}{1} \times \frac{1 \text{ mol NaOH}}{39.99g} = \frac{.125 \text{ mol}}{.750 \text{ L}} = \boxed{0.18 \text{ M}}$$

5. How many moles of Na_2CO_3 are there in 10.0 L of 2.0 M solution?

$$\frac{2.0 \text{ mol}}{1} \times \frac{10.0 \text{ L}}{1} = \boxed{20 \text{ mol Na}_2\text{CO}_3}$$

6. How many moles of Na_2CO_3 are in 10.0 mL of a 2.0 M solution?

$$\frac{2.0 \text{ mol}}{1} \times \frac{.01 \text{ L}}{1} = \boxed{0.02 \text{ mol Na}_2\text{CO}_3}$$

7. How many moles of NaCl are contained in 100.0 mL of a 0.20 M solution?

$$\frac{.20 \text{ mol}}{1} \times \frac{.100 \text{ L}}{1} = \boxed{0.02 \text{ mole NaCl}}$$

8. What weight (in grams) of NaCl would be contained in problem 7?

$$\frac{.02 \text{ mol NaCl}}{1} \times \frac{58.44 \text{ g NaCl}}{1 \text{ mol NaCl}} = \boxed{1.16 \text{ g NaCl}}$$

9. What weight (in grams) of H_2SO_4 would be needed to make 750.0 mL of 2.00 M solution?

$$\frac{2.00 \text{ mol}}{1} \times \frac{.750 \text{ L}}{1} \times \frac{98.08 \text{ g H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} = \boxed{147 \text{ g H}_2\text{SO}_4}$$

10. What volume (in mL) of 18.0 M H_2SO_4 is needed to contain 2.45 g H_2SO_4 ?

$$\frac{2.45g \text{ H}_2\text{SO}_4}{1} \times \frac{1 \text{ mol H}_2\text{SO}_4}{98.08g} \times \frac{1 \text{ L}}{18.0 \text{ mol H}_2\text{SO}_4} \times \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{1.39 \text{ mL}}$$

11. What volume (in mL) of 12.0 M HCl is needed to contain 3.00 moles of HCl?

$$\frac{3.00 \text{ mol HCl}}{1} \times \frac{1 \text{ L HCl}}{12.0 \text{ mol HCl}} \times \frac{1000 \text{ mL HCl}}{1 \text{ L HCl}} = \boxed{250 \text{ mL HCl}}$$

12. How many grams of $\text{Ca}(\text{OH})_2$ are needed to make 100.0 mL of 0.250 M solution?

$$\frac{.250 \text{ mol Ca}(\text{OH})_2}{1} \times \frac{.100 \text{ L}}{1} \times \frac{74 \text{ g Ca}(\text{OH})_2}{1 \text{ mol Ca}(\text{OH})_2} = \boxed{1.85 \text{ g Ca}(\text{OH})_2}$$

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**John A. Olmsted, Gregory M.
Williams, Robert Charles Burk**

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Chemistry John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant **Aim@AIIMS-JEE Mains** R.CHANDRAN.S.V., 2021-07-13 NEET JEE mains NCERT Based **CUET PG**

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