

Instructions, Questions, Multiple-Choice – Practice Problems

NOTE: [mg] is milligram concentration, [mg/L] is mass/volume concentration, [g/L] is volume/volume concentration, M is molarity (moles/L)

1. If 0.5g of NaCl are dissolved in 1000 mL of water, what is the % concentration (w/v) of the solution?
2. How many g of NaCl are in 500 g of solution if the concentration is 10% (w/w)?
3. A solution's concentration is 10% (v/v). How many mL of the solution do you need to have 10 mL of the solute?
4. What is the molarity (M) of a solution made by dissolving 0.10 mol of NaCl in 500 mL of water?
5. What is the molarity (M) of a solution made by dissolving 0.05 g of NaCl in 0.5 L of water? (Molar mass NaCl is 58.44 g/mol)
6. If a solution has a molarity of 1.0 M, how many L of it do you need to have 0.20 mol of the solute?
7. How many grams of MgCl_2 will you need to dissolve in 100 mL of water to create a solution with a molarity of 2.0 M? (Molar mass MgCl_2 is 95.21 g/mol)
8. You have a 1.0 M solution of concentration 0.5 M, but your patient needs a solution with a concentration of 0.2 M. How many mL of the stock solution does it take to make 500 mL of the 0.2 M solution?

Molarity And Dilution Practice Problems Key

Howard Stephen Stoker



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